

Three Hydrogen-like Atoms – Simultaneous Quantities

Electromagnetic Radiation

Consider now the hydrogen atom, which emits energy levels, light, and transitions, and also is the following quantities in terms of our understanding:

Quantum is a concept that describes the behavior of particles, such as the light, which is the electromagnetic radiation. The hydrogen atom, however, is a very simple system, which is described by quantum mechanics, and is the most

Quantity	Value (Hydrogen)	Quantum Number
Mass	Proton mass, m_p , approximately 1.67×10^{-27} kg	Proton mass, m_p , approximately 1.67×10^{-27} kg
Electron mass	Electron mass, m_e , approximately 9.11×10^{-31} kg	Electron mass, m_e , approximately 9.11×10^{-31} kg
Energy	Energy levels, E_n , approximately -13.6 eV for $n=1$	Energy levels, E_n , approximately -13.6 eV for $n=1$
Transition energy in electron volts	Transition energy, $E_{n_1 \rightarrow n_2}$, approximately 10.2 eV for $n_1=1, n_2=2$	Transition energy, $E_{n_1 \rightarrow n_2}$, approximately 10.2 eV for $n_1=1, n_2=2$
Energy of the photon (eV)	Energy of the photon, E_{ph} , approximately 10.2 eV for $n_1=1, n_2=2$	Energy of the photon, E_{ph} , approximately 10.2 eV for $n_1=1, n_2=2$

Quantum is a concept that describes the behavior of particles, such as the light, which is the electromagnetic radiation. The hydrogen atom, however, is a very simple system, which is described by quantum mechanics, and is the most

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