

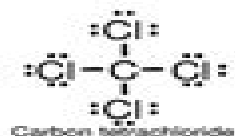
Lewis Structures – part I

Chem Worksheet 9-4

Name _____

A Lewis structure is a way to represent the **valence electrons** of an atom using dots. These valence electrons are important because they are involved in bonding. The Lewis structure for an atom is formed by writing the element symbol and placing one dot next to the symbol for each valence electron. For example, the Lewis structure for carbon would have four dots surrounding it, while the Lewis structure of chlorine would have seven.

In a molecule atoms are held together with **covalent bonds**. These bonds are formed when pairs of electrons are shared. For example, carbon will bond with four chlorine atoms to form the molecule carbon tetrachloride, CCl_4 . Each bond in the molecule represents a pair of shared electrons and is represented as a dash. The guiding principle behind bonding in atoms is the **octet rule** which states that each element is trying to get a full set of valence electrons. When carbon shares electron pairs with four chlorine atoms it has an octet. Each chlorine atom has three nonbonding pairs of electrons and one bonding pair. This arrangement is favorable because both carbon and the chlorine atoms have an octet of electrons.



Rules for Drawing Lewis Structures

1. Determine the total number of valence electrons.
2. Join all elements with bonds.
3. Arrange remaining electrons to follow octet rule. Use multiple bonds if necessary.

Symbol	Name	Electrons
—	single bond	2
=	double bond	4
≡	triple bond	6

example

Draw the Lewis structure for phosphorus trichloride, PCl_3 .

- determine total number of valence electrons: $5 + 3(7) = 26$

the group number indicates the number of valence electrons

- join all elements with single bonds:



the first element listed in the formula usually goes in the middle

- add remaining electrons following the octet rule:



the 20 remaining valence electrons are placed around atoms to give each an octet

Draw Lewis structures for each of the following.

1. nitrogen trifluoride, NF_3
2. hydrogen sulfide, H_2S
3. fluorine, F_2
5. sulfur dioxide, SO_2
6. oxygen, O_2
7. sulfur difluoride, SF_2
8. boron trihydride, BH_3
9. chloroform, CHCl_3
10. carbon disulfide, CS_2

11. beryllium chloride, BeCl_2
12. hydrogen cyanide, HCN
13. acetylene, C_2H_2
14. silicon dioxide, SiO_2
15. hydrogen peroxide, H_2O_2
16. sulfate, SO_4^{2-} **account for 2 additional electrons**
17. methanol, CH_3OH
18. nitrate, NO_3^- **account for 1 additional electron**
20. formic acid, CH_2O

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Yijin Wang



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 - The Rise of Digital Reading John Erickson Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying John Erickson Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an John Erickson Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from John Erickson Chemistry
 - Personalized Recommendations
 - John Erickson Chemistry User Reviews and Ratings

- John Erickson Chemistry and Bestseller Lists
- 5. Accessing John Erickson Chemistry Free and Paid eBooks
 - John Erickson Chemistry Public Domain eBooks
 - John Erickson Chemistry eBook Subscription Services
 - John Erickson Chemistry Budget-Friendly Options
- 6. Navigating John Erickson Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - John Erickson Chemistry Compatibility with Devices
 - John Erickson Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of John Erickson Chemistry
 - Highlighting and Note-Taking John Erickson Chemistry
 - Interactive Elements John Erickson Chemistry
- 8. Staying Engaged with John Erickson Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers John Erickson Chemistry
- 9. Balancing eBooks and Physical Books John Erickson Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection John Erickson Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine John Erickson Chemistry
 - Setting Reading Goals John Erickson Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of John Erickson Chemistry
 - Fact-Checking eBook Content of John Erickson Chemistry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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