

CHEMISTRY
Class – XII

Time – 3 hrs.

M.M. : 70

Date – 14.09.2022 (Wednesday)

Name of the student _____

Section _____

GENERAL INSTRUCTIONS:

- All questions are compulsory.
- Question nos. 1 to 21 are MCQs with one correct option and carrying 1 mark each.
- Question nos. 22 to 24 are objective/short answer type questions.
- Question nos. 25 to 28 are of 2 marks each.
- Question nos. 29 to 32 is of 3 marks each.
- Question nos. 33 to 35 is of 5 marks each.

SECTION - A

Q. 1 to 21 is of 1 mark each.

Q.1 Which of the following aqueous solution should have highest boiling point ?

- a) 1 M NaOH b) 1 M Na₂SO₄ c) 1M NH₄NO₃ d) 1 M KNO₃

Q.2 Which of the following statement is false ?

- a) Two different solution of sucrose of same molality prepare in different solvent will have the same depression in freezing point.
- b) The osmotic pressure of a solution is given by the equation $\pi = CRT$. (where C is the molarity of the solution)
- c) Decreasing order of osmotic pressure for 0.01M aqueous solution of Barium Chloride, Potassium chloride, acetic acid and sucrose is BaCl₂ > KCl > CH₃COOH > Sucrose
- d) According to Raoult's Law, the vapour pressure exerted by a volatile component of solution is directly proportional to its molar fraction in the solution.

Q.3 We have three aqueous solution of NaCl labelled as 'A', 'B' and 'C' with concentration 0.1 M, 0.01 M and 0.001 M respectively. The value of Van't Hoff factor for these solution will be in the order-

- a) $i_A < i_B < i_C$ a) $i_A > i_B > i_C$ a) $i_A = i_B = i_C$ a) $i_A < i_B > i_C$

Q.4 On the basis of the information given below mark the correct options –

- i) In bromoethane and chloroethane mixture intermolecular interaction of A-A and B-B type are nearly same as A-B type interaction.
- ii) In the ethanol and acetone mixture A-A or B-B type intermolecular interactions are stronger than A-B type interaction.

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Hussin A.Rothana



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