

Study Guide for Honors Chemistry Final - June 2015

You will be given all needed equations, values of constants and other pertinent information, including the following: periodic table, $c = \lambda\nu$, $E = h\nu$, the value for $h = 6.626 \times 10^{-34}$ J·s, the conversion between meters and nanometers ($1 \text{ m} = 1 \times 10^9 \text{ nm}$), electronegativity values, activity series of metals, gas law equations, dilution equation, molarity, molality. It is your responsibility to recognize the equations and understand how to use them. You will also be given a 3" x 5" notecard that, must be handwritten by yourself, you may use to organize your thoughts for the exam.

Do the sample problems and review all the vocabulary words in the sections indicated. Go to the following links to test your knowledge on the Chapters (remember we did not do all of the sections in all of the following chapters):

http://chemistrymc.com/chapter_test and http://chemistrymc.com/standardized_test

Electrons in Atoms – Chapter 5

Objectives:

- Describe the relationship between wavelength and frequency.
- Be able to use the equation $c = \lambda\nu$ where $c = 3.00 \times 10^8$ m/s
- Describe the relationship between frequency and energy of a wave.
- Be able to use the equation $E = h\nu$ where $h = \text{Planck's constant} = 6.626 \times 10^{-34}$ J·s
- Describe the contributions of Bohr, Schrödinger, Planck, Einstein, & deBroglie to determining the location of electrons in the atom and their properties
- Explain the difference between a continuous spectrum and a line spectrum and give examples of each.
- Be able to arrange the colors of the visible spectrum from lowest energy to highest energy.
- Define a quantum of energy and explain how it is related to an energy change of matter.
- Describe how the quantum mechanical model describes what an atom is and where the subatomic particles are located.
- Explain the relationships among an atom's energy level, sublevels, and atomic orbitals.
- Apply the Pauli Exclusion Principle, the Aufbau Principle, and Hund's Rule to write electron configurations using orbital diagrams and electron configuration notation.
- Define valence electrons and draw electron-dot structures representing an atom's valence electrons.
- Be able to read and write quantum numbers given an electron configuration or an element's symbol
- Labs – Thickness of Zinc, Spectroscopy and Flame Tests, Electron Arrangements

Vocabulary:

- | | |
|---|--|
| • Wavelength, frequency, amplitude and speed of waves | • Principal energy level |
| • Electromagnetic spectrum | • Orbitals (s, p, d, and f) |
| • Quantum | • Electron configuration |
| • Photoelectric effect | • Valence electron |
| • Atomic emission spectrum | • Electron-dot structures (Lewis dot diagrams) |
| • Ground state vs. excited state | • Quantum numbers |
| • Heisenberg uncertainty principle | • $n, \ell, m\ell, m_s$ |
| • Quantum mechanical model of the atom | |
| • Atomic orbital | |
| • Principal quantum number | |

Honors Chemistry Final Study Guide

Padhraic Smyth



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