

Section A

Answer **all** the questions in the spaces provided.

- 1 (a) The oxidation of nitrogen(II) oxide is shown in the equation.



The initial rate of this reaction was measured, starting with different concentrations of the two reactants. The following results were obtained.

experiment number	[NO] / mol dm ⁻³	[O ₂] / mol dm ⁻³	initial rate / mol dm ⁻³ s ⁻¹
1	0.032	0.012	4.08×10^{-3}
2	0.032	0.024	8.15×10^{-3}
3	0.064	0.024	3.28×10^{-2}
4	0.096	0.036	

- (i) Use the data in the table to determine the order with respect to each reactant. Show your reasoning.

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- (ii) Calculate the initial rate in experiment 4. Give your answer to **two** significant figures.

initial rate = mol dm⁻³ s⁻¹

- (iii) Write the rate equation for this reaction.

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- (iv) Use the results of experiment 1 to calculate the rate constant, k , for this reaction. Include the units of k .

rate constant, k = units

[6]

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