

# Limiting Reactants Gizmos Answer Key

The limiting reactant in a chemical reaction is the reactant that

- A. is used up at the end of the reaction.
  - B. is left over after the reaction.
  - C. has the largest number of atoms.
  - D. has the fewest number of atoms. - correct answer
- used up at the end of the reaction.

Correct Answer: A. is

How many water molecules (H<sub>2</sub>O) can be produced from 6 molecules of hydrogen gas (white) reacting with 6 molecules of oxygen gas (red)? Which reactant is the limiting reactant?

- A. 6 water molecules, hydrogen is limiting
  - B. 3 water molecules, oxygen is limiting
  - C. 3 water molecules, hydrogen is limiting
  - D. 12 water molecules, oxygen is limiting - correct answer
- 6 water molecules, hydrogen is limiting

Correct Answer: A.

34 atoms of carbon (C) react with 22 molecules of hydrogen gas (H<sub>2</sub>). How many molecules of methane (CH<sub>4</sub>) will be formed, and what will be left over?

- A. 17 molecules of methane formed, 5 molecules of hydrogen left over
  - B. 11 molecules of methane formed, 23 molecules of carbon left over
  - C. 5 molecules of methane formed, 24 atoms of carbon and 2 molecules of hydrogen left over
  - D. 22 molecules of methane formed, 12 atoms of carbon left over - correct answer
- Correct Answer: B. 11 molecules of methane formed, 23 molecules of carbon left over

When aluminum (Al) reacts with chlorine gas (Cl<sub>2</sub>), aluminum chloride (AlCl<sub>3</sub>) is produced according to the balanced equation below:  $2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$  If 20 atoms of aluminum react with 45 molecules of chlorine gas, which reactant is limiting and how many more atoms/molecules would be required to use up all the reactants?

- A. Chlorine is limiting, 5 more molecules are needed.

# Gizmo Limiting Reactants Answer

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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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