

10. Two ideal gases X and Y are at the same temperature. The mass of the molecules of gas X is twice the mass of the molecules of gas Y.

What is the ratio $\frac{\text{average speed of molecules of gas X}}{\text{average speed of molecules of gas Y}}$?

A. $\frac{1}{2}$

Microscopic level: (absolute) **temperature** directly proportional to the **average** kinetic energy of the molecules of a substance:

B. $\frac{1}{\sqrt{2}}$

B

$$\langle KE \rangle_{avg} = \frac{3}{2} kT \quad k \text{ is Boltzmann constant}$$

C. $\sqrt{2}$

$$\frac{1}{2} m \langle v^2 \rangle_{avg} \sim T$$

D. 2

$$\frac{\frac{1}{2} m_x \langle v^2 \rangle_{x avg}}{\frac{1}{2} m_y \langle v^2 \rangle_{y avg}} = \frac{T}{T}$$

$$\frac{\langle v^2 \rangle_{x avg}}{\langle v^2 \rangle_{y avg}} = \frac{m_y}{m_x} = \frac{1}{2}$$

M13 4 Physi Sp1 Tz

Thirumalaisamy P.
Velavancorresponding



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- Chapter 5: Conclusion

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