

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

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Bud Morgan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the Bud Morgan text.

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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures Prentice Hall Science Explorer Michael J. Padilla, 2002 The Science Teachers Bulletin ,1989

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Nuclear Science Abstracts** ,1974

Chemistry John S. Phillips, Cheryl Wistrom, 2000 *Chemistry3* Andrew Burrows, John Holman, Andrew Parsons, Gwen Pilling, Gareth Price, 2017 *Chemistry3* establishes the fundamental principles of all three strands of chemistry organic inorganic and physical By building on what students have learned at school using carefully worded explanations annotated diagrams and worked examples it presents an approachable introduction to chemistry and its relevance to everyday life

Fusion Energy Update ,1985 **Physics Briefs** ,1994 Chemistry Leonard W. Fine, 1972 **The Trend in Engineering at the University of Washington** ,1961 **Government Reports Announcements & Index** ,1983

INIS Atomindex ,1987 **Ceramic Abstracts** ,1953 Japanese Technical Abstracts ,1987 Electrical & Electronics Abstracts ,1995 **International Aerospace Abstracts** ,1998 **Development of a Periodic Function for Atomic Radius** Jaswinder Kaur, Jaspreet Kaur, 2011-12

A number of experimental methods had been developed till date to evaluate and study the periodic properties To study these periodic properties theoretically we have developed a new linear approach A mathematical periodic function has been developed to calculate atomic radii covalent and vander Waal of any element by knowing just its atomic number This function consists of a common equation along with additional terms correction factor for different blocks s p d and f of periodic table This function has been derived mathematically by relating different factors in form of correction factors which affect the atomic radii using graphs It takes into accounts the chemical behaviour of all the elements in their respective groups and periods *Explanatory Key to the Periodic Chart of the Atoms* Henry D. Hubbard, 1926 ATOMIC STRUCTURE High School Chemistry Curriculum Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons

within an atom's electron cloud including energy levels, sublevels, and orbital shapes. We will also discuss the principles of electron configuration and how it impacts an atom's chemical properties. Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure. We will explore periodic trends such as atomic radius, ionization energy, and electron affinity, and discuss how these trends are related to an element's position on the table. In our final chapter, we will discuss the significance of atomic structure in various chemical reactions, including bonding, stoichiometry, and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry, environmental science, and materials science. Throughout this book, we will provide clear explanations, helpful illustrations, and interactive activities to reinforce key concepts. Whether you are a visual learner, a hands-on learner, or a theoretical learner, this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

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