

**5 (d)** Explain why the current in the external circuit of this cell falls to zero after the cell has operated for some time.

.....

.....

*(1 mark)*

**5 (e)** The simplified electrode reactions in a rechargeable lithium cell are

Electrode **A**       $\text{Li}^+ + \text{MnO}_2 + \text{e}^- \longrightarrow \text{LiMnO}_2$        $E = -0.15 \text{ V}$

Electrode **B**               $\text{Li}^+ + \text{e}^- \longrightarrow \text{Li}$

Electrode **B** is the negative electrode.

**5 (e) (i)** The e.m.f. of this cell is 2.90 V.

Use this information to calculate a value for the electrode potential of electrode **B**.

.....

.....

*(1 mark)*

**5 (e) (ii)** Write an equation for the overall reaction that occurs when this lithium cell is being recharged.

.....

.....

.....

*(2 marks)*

**5 (e) (iii)** Suggest why the recharging of a lithium cell may lead to release of carbon dioxide into the atmosphere.

.....

.....

*(1 mark)*

# June 2013 Chemistry 97paper 35 Question

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