

Section 7.2: Mass Relationships in Chemical Reactions

Tutorial 1 Practice, page 323–324

1. (a) Given: $m_{\text{CaCO}_3} = 0.50 \text{ g}$

Required: mass of hydrochloric acid, m_{HCl}

Solution:

Step 1. List the given value, the required value, and the corresponding molar masses.



0.50 g m_{HCl}
100.09 g/mol 36.46 g/mol

Step 2. Convert mass of calcium carbonate to amount of calcium carbonate.

$$n_{\text{CaCO}_3} = 0.50 \text{ g} \times \frac{1 \text{ mol}_{\text{CaCO}_3}}{100.09 \text{ g}}$$

$$n_{\text{CaCO}_3} = 0.004\,996 \text{ mol} \quad [\text{two extra digits carried}]$$

Step 3. Convert amount of calcium carbonate to amount of hydrochloric acid.

$$n_{\text{HCl}} = 0.004\,996 \text{ mol}_{\text{CaCO}_3} \times \frac{2 \text{ mol}_{\text{HCl}}}{1 \text{ mol}_{\text{CaCO}_3}}$$

$$n_{\text{HCl}} = 0.009\,992 \text{ mol} \quad [\text{two extra digits carried}]$$

Step 4. Convert amount of hydrochloric acid to mass of hydrochloric acid.

$$m_{\text{HCl}} = (0.009\,992 \text{ mol}) \left(\frac{36.46 \text{ g}}{1 \text{ mol}} \right)$$

$$m_{\text{HCl}} = 0.36 \text{ g}$$

Statement: 0.50 g of calcium carbonate will neutralize 0.36 g of hydrochloric acid.

(b) Given: $m_{\text{CaCO}_3} = 0.50 \text{ g}$; $n_{\text{CaCO}_3} = 0.004\,996 \text{ mol}$

Required: mass of calcium chloride, m_{CaCl_2}

Solution:

Step 1. List the given values, the required values, and the corresponding molar masses.



0.004 996 mol n_{CaCl_2}
0.50 g m_{CaCl_2}
100.09 g/mol 110.98 g/mol

Step 2. Convert amount of calcium carbonate to amount of calcium chloride.

$$n_{\text{CaCl}_2} = n_{\text{CaCO}_3}$$

$$n_{\text{CaCl}_2} = 0.004\,996 \text{ mol}$$

Mass Relationships In Chemical Reactions Practice Test

Marcel A. Müller



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Table of Contents Mass Relationships In Chemical Reactions Practice Test

1. Understanding the eBook Mass Relationships In Chemical Reactions Practice Test
 - The Rise of Digital Reading Mass Relationships In Chemical Reactions Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Mass Relationships In Chemical Reactions Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mass Relationships In Chemical Reactions Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mass Relationships In Chemical Reactions Practice Test
 - Personalized Recommendations
 - Mass Relationships In Chemical Reactions Practice Test User Reviews and Ratings
 - Mass Relationships In Chemical Reactions Practice Test and Bestseller Lists

5. Accessing Mass Relationships In Chemical Reactions Practice Test Free and Paid eBooks
 - Mass Relationships In Chemical Reactions Practice Test Public Domain eBooks
 - Mass Relationships In Chemical Reactions Practice Test eBook Subscription Services
 - Mass Relationships In Chemical Reactions Practice Test Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Mass Relationships In Chemical Reactions Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mass Relationships In Chemical Reactions Practice Test
 - Highlighting and Note-Taking Mass Relationships In Chemical Reactions Practice Test
 - Interactive Elements Mass Relationships In Chemical Reactions Practice Test
8. Staying Engaged with Mass Relationships In Chemical Reactions Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mass Relationships In Chemical Reactions Practice Test
9. Balancing eBooks and Physical Books Mass Relationships In Chemical Reactions Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mass Relationships In Chemical Reactions Practice Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mass Relationships In Chemical Reactions Practice Test
 - Setting Reading Goals Mass Relationships In Chemical Reactions Practice Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mass Relationships In Chemical Reactions Practice Test
 - Fact-Checking eBook Content of Mass Relationships In Chemical Reactions Practice Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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