

# MASS MOLE

## STOICHIOMETRY WORKSHEET

1. Consider the following reaction:



- If 15 g of  $\text{N}_2\text{O}_4$  was produced, how many moles of  $\text{O}_2$  were required?
- If  $4 \times 10^{-2}$  moles of oxygen reacted, how many grams of  $\text{N}_2$  were needed?

2. Consider the following reaction:



- What is the mass of potassium nitrate produced when 2.04 moles of potassium phosphate react?
- If 5.8 g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

3. Consider the following reaction:



- If you have 5.5 mol of  $\text{CaC}_2$ , how much  $\text{C}_2\text{H}_2$  do you get?
- How many moles of water are needed when 65 g of  $\text{CaC}_2$  have reacted?

4. Hydrogen sulfide gas burns in oxygen to produce sulfur dioxide and water vapor:



What mass of oxygen gas is consumed in a reaction that produces 4.6 mol  $\text{SO}_2$ ?

# Mass Mass Stoichiometry Packet Answers

**Lingsheng Yao**



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