

4. Which of the following substances would have hydrogen bonding as one of its IMF's?



5. Which molecular substance from the following list would have the highest boiling point?

- A. Ne
- B. F_2
- C. H_2
- D. Cl_2

6. Often, the actual gas pressure is less than the pressure predicted by the ideal gas law; the explanation lies in the fact that the ideal gas law does NOT include a factor for molecular

- A. volume
- B. mass
- C. velocity
- D. attraction
- E. shape

7. 25.0 L of hydrogen gas is collected over water at 40°C . The total pressure of the sample is 750 mmHg. At 40°C , the vapor pressure of water is 55 mmHg. What is the mass of the hydrogen gas?

- A. 1.00 g
- B. 1.50 g
- C. 0.0060 g
- D. 0.0607 g
- E. 1.12 g

8. A 2.00 L sample of nitrogen gas at 27°C and 600 mmHg is heated until it occupies a volume of 3.00 L. If the pressure remains unchanged, the final temperature of the gas is

- A. 40°C
- B. 120°C
- C. 477°C
- D. 677°C
- E. 900°C

9. PI (d) bonding occurs in each of the following EXCEPT

- A. C_2H_2
- B. C_2H_4
- C. C_2H_6
- D. C_2H_2
- E. C_2H_4

10. Determine the shape and polarity of the PF_3 molecule

- A. trigonal pyramidal, polar
- B. trigonal pyramidal, nonpolar
- C. trigonal planar, polar
- D. trigonal planar, nonpolar

High School Chemistry Final Exam Answers

Rachel S Tattersall



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