

2

MATTER AND CHANGE

SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The _____ mass _____ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a _____ substance _____.

4. Is the following sentence true or false? All samples of a substance have different physical properties. _____ false _____

5. A physical property is a quality or condition of a substance that can be _____ observed _____ or _____ measured _____ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness c. boiling point
b. color d. melting

7. Look at Table 2.1 on page 40. What is the melting point of bromine? _____ -7°C _____

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.

- a. sulfur
b. chlorine
c. gold
d. copper

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Richard J. D. Tilley



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Table of Contents Matter And Change Workbook Pearson

1. Understanding the eBook Matter And Change Workbook Pearson
 - The Rise of Digital Reading Matter And Change Workbook Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Matter And Change Workbook Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Matter And Change Workbook Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Matter And Change Workbook Pearson
 - Personalized Recommendations
 - Matter And Change Workbook Pearson User Reviews and Ratings
 - Matter And Change Workbook Pearson and Bestseller Lists
5. Accessing Matter And Change Workbook Pearson Free and Paid eBooks
 - Matter And Change Workbook Pearson Public Domain eBooks
 - Matter And Change Workbook Pearson eBook Subscription Services
 - Matter And Change Workbook Pearson Budget-Friendly Options

6. Navigating Matter And Change Workbook Pearson eBook Formats
 - ePub, PDF, MOBI, and More
 - Matter And Change Workbook Pearson Compatibility with Devices
 - Matter And Change Workbook Pearson Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Matter And Change Workbook Pearson
 - Highlighting and Note-Taking Matter And Change Workbook Pearson
 - Interactive Elements Matter And Change Workbook Pearson
8. Staying Engaged with Matter And Change Workbook Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Matter And Change Workbook Pearson
9. Balancing eBooks and Physical Books Matter And Change Workbook Pearson
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Matter And Change Workbook Pearson
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Matter And Change Workbook Pearson
 - Setting Reading Goals Matter And Change Workbook Pearson
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Matter And Change Workbook Pearson
 - Fact-Checking eBook Content of Matter And Change Workbook Pearson
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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