

**JEE 2009  
Paper I**

**PART I : CHEMISTRY**

**SECTION - I**

**Single Correct Choice Type**

This section contains 8 multiple choice questions. Each question has 4 choices (A), (B), (C) and (D) for its answer, out of which **ONLY ONE** is correct.

1. Given that the abundances of isotopes  $^{54}\text{Fe}$ ,  $^{56}\text{Fe}$  and  $^{58}\text{Fe}$  are 5%, 90% and 5%, respectively, the atomic mass of Fe is  
(A) 55.85 (B) 55.95 (C) 55.75 (D) 56.05

**Answer**

- (A) ☐ (B) ☒ (C) ☐ (D) ☐

2. The term that corrects for the attractive forces present in a real gas in the van der Waals equation is

- (A)  $ab$  (B)  $\frac{an^2}{V^2}$  (C)  $-\frac{an^2}{V^2}$  (D)  $-ab$

**Answer**

- (A) ☐ (B) ☒ (C) ☐ (D) ☐

3. Among the electrolytes  $\text{Na}_2\text{SO}_4$ ,  $\text{CaCl}_2$ ,  $\text{Al}_2(\text{SO}_4)_3$  and  $\text{NH}_4\text{Cl}$ , the most effective coagulating agent for  $\text{Sb}_2\text{S}_3$  sol is

- (A)  $\text{Na}_2\text{SO}_4$  (B)  $\text{CaCl}_2$  (C)  $\text{Al}_2(\text{SO}_4)_3$  (D)  $\text{NH}_4\text{Cl}$

**Answer**

- (A) ☐ (B) ☐ (C) ☒ (D) ☐

4. The Henry's law constant for the solubility of  $\text{N}_2$  gas in water at 298 K is  $1.0 \times 10^5$  atm. The mole fraction of  $\text{N}_2$  in air is 0.8. The number of moles of  $\text{N}_2$  from air dissolved in 10 moles of water at 298 K and 5 atm pressure is

- (A)  $4.0 \times 10^{-4}$  (B)  $4.0 \times 10^{-5}$  (C)  $5.0 \times 10^{-4}$  (D)  $4.0 \times 10^{-6}$

**Answer**

- (A) ☒ (B) ☐ (C) ☐ (D) ☐

5. The reaction of  $\text{P}_4$  with X leads selectively to  $\text{P}_4\text{O}_6$ . The X is

- (A) Dry  $\text{O}_2$  (B) A mixture of  $\text{O}_2$  and  $\text{N}_2$   
(C) Moist  $\text{O}_2$  (D)  $\text{O}_2$  in the presence of aqueous  $\text{NaOH}$

**Answer**

- (A) ☐ (B) ☒ (C) ☐ (D) ☐

# Iit Exam Paper 2009

**B Lingard**



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