

**Solution**  
**CHAPTER TEST**  
**Class 11 - Chemistry**

1. (a) Propene

**Explanation:** Elimination reaction takes place, a molecule of HBr gets eliminated to form propene.

2. (c)  $\text{HI} > \text{HBr} > \text{HCl}$

**Explanation:**  $\text{HI} > \text{HBr} > \text{HCl}$

3. (a) 1,3-dimethylbenzene



4. (d) Geometric isomers

**Explanation:** Isomers which have the same structural formulae but differ in the relative spatial arrangements of atoms or groups around the double bond are called geometric isomers. Thus geometrical isomers are a type of stereoisomers.

5. (d)  $\text{HgSO}_4$

**Explanation:** Alkynes are not readily hydrated due to their low reactivity towards electrophilic addition reactions. Further dilute  $\text{H}_2\text{SO}_4$  has no effect on carbon carbon triple bond. In the presence of  $\text{HgSO}_4$  as catalyst, hydration occurs readily.

6. (a) A mixture of low boiling hydrocarbons.

**Explanation:** Pyrolysis is the decomposition of a compound by the application of heat. It involves the breaking of C-C and C-H bonds resulting in the formation of lower boiling alkanes.

7. (a) 3

**Explanation:** The position isomers of  $\text{C}_5\text{H}_8$  are n-pentyne, 3-methyl butyne, pent-2-yne.

8. (a)  $\text{C}_8\text{H}_{18}$

**Explanation:** Alkenes are unsaturated hydrocarbons with at least one double bond between the carbon atoms. Their general formula is  $\text{C}_n\text{H}_{2n}$ .

9. (b) Sodium acetate

**Explanation:**

This is an example of Kolbe's electrolysis method. The reaction is:



The step-wise redox reactions occurring in the electrolytic cell are depicted as under



At Anode: (Oxidation)



At Cathode: (Reduction)



# Hydrocarbons Answer Key

**LL Leslie**

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## **Table of Contents Hydrocarbons Answer Key**

1. Understanding the eBook Hydrocarbons Answer Key
  - The Rise of Digital Reading Hydrocarbons Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Hydrocarbons Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hydrocarbons Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hydrocarbons Answer Key
  - Personalized Recommendations
  - Hydrocarbons Answer Key User Reviews and Ratings
  - Hydrocarbons Answer Key and Bestseller Lists
5. Accessing Hydrocarbons Answer Key Free and Paid eBooks
  - Hydrocarbons Answer Key Public Domain eBooks
  - Hydrocarbons Answer Key eBook Subscription Services
  - Hydrocarbons Answer Key Budget-Friendly Options
6. Navigating Hydrocarbons Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Hydrocarbons Answer Key Compatibility with Devices
- Hydrocarbons Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hydrocarbons Answer Key
  - Highlighting and Note-Taking Hydrocarbons Answer Key
  - Interactive Elements Hydrocarbons Answer Key
- 8. Staying Engaged with Hydrocarbons Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hydrocarbons Answer Key
- 9. Balancing eBooks and Physical Books Hydrocarbons Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hydrocarbons Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hydrocarbons Answer Key
  - Setting Reading Goals Hydrocarbons Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hydrocarbons Answer Key
  - Fact-Checking eBook Content of Hydrocarbons Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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