

1st Semester Chemistry Final Exam Review

1. Identify the following reaction type AND correct products:



- a. combustion
- b. synthesis
- c. decomposition
- d. single replacement
- e. double replacement

- w. $\text{Ag} + \text{ScOH}$
- x. $\text{Ag} + 2 \text{Sc}(\text{OH})_2$
- y. $3 \text{Ag} + \text{Sc}(\text{OH})_3$
- z. $3 \text{Ag} + 2 \text{Sc}(\text{OH})_3$

- | | |
|-------|---------|
| I. | a. & w. |
| II. | b. & z. |
| III. | b. & y. |
| IV. | d. & z. |
| V. | e. & z. |
| VI. | d. & y. |
| VII. | e. & y. |
| VIII. | c. & z. |
| IX. | d. & x. |

2. Convert: $2.32 \times 10^5 \text{ mg} = ? \text{ kg}$

3. Calculate the energy for a quantum of light with a frequency of $8.44 \times 10^{18} \text{ Hz}$.

4. Determine the oxidation number for iodine in aluminum iodate.

5. REPORTING WITH THE CORRECT NUMBER OF SIGNIFICANT DIGITS, how many grams are in **4.011 moles** of calcium chloride?

- a) 445.14078 grams b) 445.140 grams c) 445.1 grams d) 445 grams
e) 450 grams f) 400 grams

6. Calculate the frequency of light that has a wavelength of 210 nm.

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RS Peters

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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