

100

27. It is necessary to know the formula of a compound in order to calculate its percentage composition.

AT

28. If the percent by mass of carbon in methane, CH_4 , is 75%, then 100 grams of methane contains 75 g of carbon and 25 g of hydrogen.

AT

29. The formula for methane, CH_4 , is both a molecular and an empirical formula.

100

30. The empirical formula of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, is CH_2O .

C. Problems

Solve the following problems on the space provided. Show your work.

31. How many moles of Fe are in 5.1 g of 100% pure Fe?

32. How many representative particles are in 2.11 g of 100% pure MgCl_2 ?

33. What is the gram molecular mass of carbon, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$?

34. Find the mass of each of the following:

a. 2.43 g of 100% pure H_2O_2

b. 2.43 g of H_2O_2

c. 4.86 g of 100% pure H_2O_2

d. 4.86 g of H_2O_2

Honors Chemistry Chemical Quantities Answer Key

CH Cherryholmes



Honors Chemistry Chemical Quantities Answer Key:

Honors Chemistry Chemical Quantities Answer Key Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the power of words has been evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Honors Chemistry Chemical Quantities Answer Key**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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