

Equilibria Involving A Weak Base

You have 0.010 M NH_3 . Calc. the pH.



$$K_b = 1.8 \times 10^{-5}$$

Step 1. Define equilibrium concs. in ICE table

	$[\text{NH}_3]$	$[\text{NH}_4^+]$	$[\text{OH}^-]$
initial	0.010	0	0
change	-x	+x	+x
equilib	0.010 - x	x	x

Honors Chemistry Equilibrium

ML Morrison



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Principles of Modern Chemistry David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Fast**

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Edition is a fully revised and updated textbook that fills the need for a comprehensive treatment of aquatic chemistry and covers the many complicated equations and principles of aquatic chemistry It presents the established science of equilibrium water chemistry using the uniquely recognizable step by step Pankow format which allows a broad and deep understanding of aquatic chemistry The text is appropriate for a wide audience including undergraduate and graduate students industry professionals consultants and regulators Every professional using water chemistry will want this text within close reach and students and professionals alike will expect to find at least one copy on their library shelves Key Features Extremely thorough one of a kind treatment of aquatic chemistry which considers a chemical thermodynamics fundamentals b acid base titration and buffer calculations c CO₂ chemistry and alkalinity d complexation of metal ions by ligands and chelates e mineral solubility processes f redox chemistry including the chemistry of chlorine as in disinfection oxygen CO₂ and methane nitrogen sulfur iron and lead including the story of lead in the drinking water of Flint Michigan and g electrical effects in aqueous solutions including the Debye H ckel Law and related equations for activity corrections double layers and colloid

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University of Michigan Official Publication , 1959 Catalogue of the University of Michigan University of Michigan, 1967 Announcements for the following year included in some vols **Announcement** University of Michigan. College of Engineering, 1967 *Text Book of Physical Chemistry Semester-IV, Course-10* by BVR VENKATA RAO BASA (BVR SIR), 2025-03-12 Text Book of Physical Chemistry Semester IV Course 10 by BVR is meant for BSc students pursuing Organic Chemistry Honors The syllabus may vary slightly from college to college The topics covered are dilute solutions electrochemistry phase rule photochemistry Practicals CST and conductometry Objective questions for entrance exams included and also video links for further knowledge included Bulletin of Information United States Coast Guard Academy, 1976 **Catalogue** Princeton University, 1917

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