

1. The theory that explains the behavior of gases at the molecular level is called the kinetic theory. It states that the molecules of a gas are in constant motion and that the pressure exerted by the gas is due to the collisions of the molecules with the walls of the container.
2. Gases behave much like ideal gas behavior under conditions of very low pressure and very high temperature. The molecules of an ideal gas display no attraction or repulsion for one another. Under ordinary conditions, an ideal gas consists chiefly of single atoms which combine only weakly and are easily separated. Ideal gas particles move in straight lines until they collide with each other or with the walls of their container. The average kinetic energy of the molecules of an ideal gas is proportional to the square of the absolute temperature of the gas.
3. A gas exerts pressure on the walls of its container because gas molecules collide with the walls of the container. As temperature increases, gas molecules move faster. As the temperature of the gas increases, the average speed of the molecules increases. The pressure of the gas increases because the rate of the collisions increases.
4. To measure gas pressure an instrument called a manometer is used.
5. The earth's atmosphere has weight, which creates atmospheric pressure.
6. The instrument used to measure atmospheric pressure is the barometer.

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