

- 1** This question is about reversible reactions.

Sulfur trioxide is produced in a reversible reaction.

The equation for the reaction is:



- 1.1** What is the correct balanced equation for the reaction?

Tick **one** box.

- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
☐ $\text{SO}_2 + 2\text{O} \rightleftharpoons \text{SO}_3$
☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$
☐ $2\text{SO}_2 + 2\text{O}_2 \rightleftharpoons 2\text{SO}_3$

[1 mark]

- 1.2** The reaction of sulfur dioxide and oxygen to form sulfur trioxide is exothermic.

Which of the following statements is true?

Tick **one** box.

- ☐ The activation energy for the reaction of sulfur dioxide with oxygen is negative.
☐ Sulfur trioxide is colder than sulfur dioxide and oxygen.
☐ Heating the reaction will cause more sulfur trioxide to form.
☐ The reaction of sulfur trioxide to form sulfur dioxide and oxygen is endothermic.

[1 mark]

Exam Chemistry June 2007 Paper 6 Cambridge

K Morrison



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