

change of the surrounding water.  $\Delta H_{\text{metal}} = -\Delta H_{\text{water}}$  If the water temperature increases, then the metal must have released heat. We can find the enthalpy change of the metal with the formula:  $\Delta H_{\text{metal}} = -m_{\text{water}} \cdot C_{\text{water}} \cdot \Delta T_{\text{water}}$

Check-for-understanding question:

1. When 25 mL of water containing 0.85 mol HCl is mixed with 40 mL of water containing .05 mol KOH in a calorimeter (with initial temperature 23 °C), the temperature of the solution increases to 25.5 °C. How many kilojoules (kJ) of heat are released by the acid-base reaction?

$$\begin{aligned}\Delta H_{\text{water}} &= m_{\text{water}} \times C_{\text{water}} \times \Delta T_{\text{water}} \\ &= 65 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 2.5^\circ\text{C} \\ &= 679.25 \text{ J} = .67925 \text{ kJ}\end{aligned}$$

# Measuring And Expressing Enthalpy Changes Answers

**John Newman, Karen E. Thomas-Alyea**



## Measuring And Expressing Enthalpy Changes Answers:

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**Physical Chemistry for the Life Sciences** Peter William Atkins,R. George Ratcliffe,Julio de Paula,Mark Wormald,2023 This text provides a balanced presentation of the concepts of physical chemistry and their applications to biology and biochemistry Written to straddle the worlds of physical chemistry and the life sciences it shows students how the tools of physical chemistry can elucidate biological questions

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