



# Symbol LS2208

## Product Reference Guide



# Ls2208 Product Reference Guide

**H Kauffman**



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web the balanced chemical equation for this reaction is  $\text{Mg}(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{Mg}(\text{NO}_3)_2(\text{aq}) + \text{H}_2(\text{g})$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

$\text{HNO}_3 + \text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2 \rightarrow \text{H}_2\text{C}_2\text{H}_3\text{O}_2 + \text{Mg}(\text{NO}_3)_2$  chemical equation - Aug 11 2023

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web mar 4 2018  $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$  on the left side there are 1 magnesium in elemental form  $\text{H}_2$  there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6

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[MgCO<sub>3</sub> + HNO<sub>3</sub> → Mg\(NO<sub>3</sub>\)<sub>2</sub> + H<sub>2</sub>O + CO<sub>2</sub> chemical equation](#) - Jul 10 2023

web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate  $\text{MgCO}_3$  and two moles of nitric acid  $\text{HNO}_3$  react to form one mole of magnesium nitrate  $\text{Mg}(\text{NO}_3)_2$  one mole of water  $\text{H}_2\text{O}$  and one mole of carbon dioxide  $\text{CO}_2$  show chemical structure image reactants

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[Mg\(NO<sub>3</sub>\)<sub>2</sub> + H<sub>2</sub>SO<sub>4</sub> → MgSO<sub>4</sub> + HNO<sub>3</sub> chemical equation](#) - Apr 07 2023

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