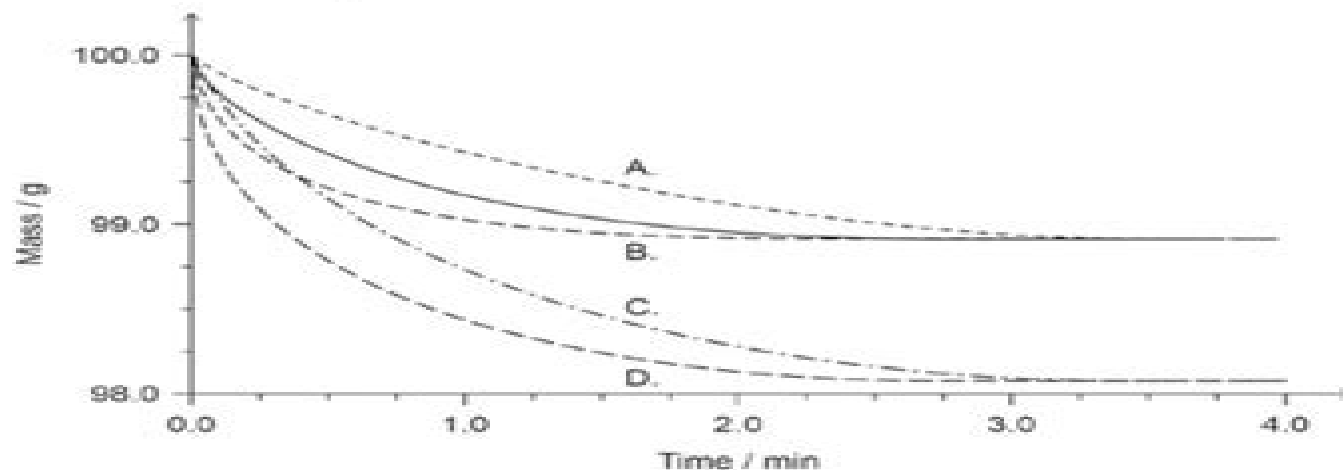


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1. [1 mark]

A sample of calcium carbonate reacts with excess hydrochloric acid in a beaker. The solid line shows how the mass of the beaker changes with time.

Which dashed line represents the results obtained when the acid concentration is doubled?



2. [1 mark]

Why does a reaction for a sample of gases, at constant temperature, occur faster at higher pressure?

- A. Collisions are more frequent.
- B. Collisions are more energetic.
- C. High pressure lowers activation energy.
- D. The reaction is more exothermic at high pressure.

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