

4.7 CHEMISTRY (20%)

4.7.1 Chemistry Paper 4 (20%)

4. (a) Ethene (1 mark)
- (b) Add water to each compound.
Add NaOH or NaHCO_3 to a suitable salt of each compound.
- $\text{C}_2\text{H}_4\text{O}$ - no effervescence (1 mark)
- $\text{C}_2\text{H}_4\text{O}_2$ effervescence (1 mark)
- no eff. with either potassium dichromate (VII) and water
 $\text{C}_2\text{H}_4\text{O}$ - colour from orange to green
- $\text{C}_2\text{H}_4\text{O}_2$ - no change
no eff. with either Potassium Manganate (VII)
 $\text{C}_2\text{H}_4\text{O}$ - with $\text{Fe}(\text{SO}_4)_2$
- $\text{C}_2\text{H}_4\text{O}_2$ - no change
5. (a) Ethane (2 marks)
- (b) — Sodium is very reactive than electrolysis (1 mark)
— More reactive than carbon (1 mark)
- (c) Ethane (1 mark)
- Sodium lamp, used in nuclear reactions
Sodium vapour, sodium amalgam
 NaOH , Electrolysis and chlorates, etc.
6. (a) Ethylalco - change, white one side of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Ethylalcoholic solution (percent) (1 mark)
7. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure at the product of pressure and volume is constant at a fixed temperature. (2 marks)

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N Colangelo



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Table of Contents January 2015 Chemistry Igcse Mark Scheme

1. Understanding the eBook January 2015 Chemistry Igcse Mark Scheme
 - The Rise of Digital Reading January 2015 Chemistry Igcse Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying January 2015 Chemistry Igcse Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an January 2015 Chemistry Igcse Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from January 2015 Chemistry Igcse Mark Scheme
 - Personalized Recommendations
 - January 2015 Chemistry Igcse Mark Scheme User Reviews and Ratings
 - January 2015 Chemistry Igcse Mark Scheme and Bestseller Lists

5. Accessing January 2015 Chemistry Igcse Mark Scheme Free and Paid eBooks
 - January 2015 Chemistry Igcse Mark Scheme Public Domain eBooks
 - January 2015 Chemistry Igcse Mark Scheme eBook Subscription Services
 - January 2015 Chemistry Igcse Mark Scheme Budget-Friendly Options
6. Navigating January 2015 Chemistry Igcse Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - January 2015 Chemistry Igcse Mark Scheme Compatibility with Devices
 - January 2015 Chemistry Igcse Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of January 2015 Chemistry Igcse Mark Scheme
 - Highlighting and Note-Taking January 2015 Chemistry Igcse Mark Scheme
 - Interactive Elements January 2015 Chemistry Igcse Mark Scheme
8. Staying Engaged with January 2015 Chemistry Igcse Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers January 2015 Chemistry Igcse Mark Scheme
9. Balancing eBooks and Physical Books January 2015 Chemistry Igcse Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection January 2015 Chemistry Igcse Mark Scheme
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine January 2015 Chemistry Igcse Mark Scheme
 - Setting Reading Goals January 2015 Chemistry Igcse Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of January 2015 Chemistry Igcse Mark Scheme
 - Fact-Checking eBook Content of January 2015 Chemistry Igcse Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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