

GENERAL INORGANIC CHEMISTRY

Study of Matter

A. Classification of Matter

1. Pure Substance

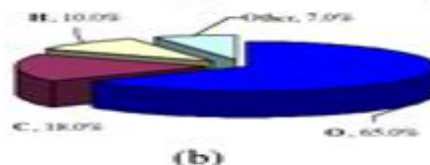
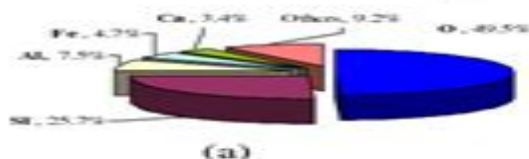
- form of matter that has distinct composition and properties

a. Elements

- Substance composed of one kind of atom and cannot be decomposed into simpler substances
- There are 118 known elements of which 91 elements exist in the Earth in macroscopic or trace quantities, 3 elements (Tc, Pm and Np) exist as a result of rare decay of heavy elements and 24 elements (Am to Uuo) are synthetically prepared

b. Compounds

- Substance formed when elements interact with other elements
- A pure compound is made up of the elements in the same proportion by mass – Law of Definite Proportions by Joseph Louis Proust (1754-1826)
- Abundance of the elements in (a) earth's crust and (b) human body



2. Mixture

- form of matter that consists of different substances and whose composition varies
- a. **Heterogeneous mixture** – mixture that does not have uniform composition, properties and appearance
- b. **Homogeneous mixture** – mixture that is uniform throughout

B. Properties of Matter

1. **Physical properties** – observed without changing the identity and composition of the substance
2. **Chemical properties** – observed when a substance reacts to form other substances
3. **Intensive properties** – properties independent of the amount of substance present such as melting point, density, temperature, etc.
4. **Extensive properties** – properties dependent of the amount of substance present such as mass and volume

C. Changes in Matter

1. **Physical change** – changes that do not result in the formation of new substances
2. **Chemical change** – changes that result in the formation of new substances

Atomic Theory

A. Dalton's Atomic Theory

- Proposed by John Dalton (1766-1844)

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Mary Burnham, Carol Hurd



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