

Practice Test #15: Chemical Equilibrium

1. Which one of the following will change the value of an equilibrium constant?

- A) Changing the temperature
- B) Adding other substances that do not react with any of the species involved in the equilibrium
- C) Varying the initial concentrations of reactants
- D) Varying the initial concentrations of products

2. An aqueous equilibrium mixture of CoCl_4^{2-} , CoBr_4^{2-} , Cl^- , and Br^- is present in a flask at 25°C . Which action below, will change the value of the equilibrium constant from that which currently describes the concentration relationships of the four species above:

- A) add more Cl^- to the solution
- B) add more Br^- to the solution
- C) add more CoBr_4^{2-} to the solution
- D) add more CoCl_4^{2-} to the solution
- E) put the flask into an 80°C water bath

3. The equilibrium constant for reaction (1) is "K". What is the equilibrium constant for equation (2)?



- A) K^2
- B) $2K$
- C) $\frac{1}{2} K$
- D) $1/K^2$

The second equation is both flipped and multiplied by 2, resulting in the value of K being inverted and squared.

4. The value of K for the following reaction is 0.25.



What is the value of K for the reaction shown below?



- A) 0.50
- B) 0.0625
- C) 0.12
- D) 0.25

The second equation is the same as the first except that it is multiplied by 2.

The new value of K is $0.25^2 = 0.0625$

Holt Chemistry Chemical Equilibrium Answer Key Test

John Victor Walther



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