

Topic: Photosynthesis

Activity 1

Give the correct biological term for each of the following descriptions.

No.	Description	Biological Term
1.1	A molecule that serves as an energy carrier during photosynthesis to help make organic molecules.	
1.2	Openings in a leaf through which gaseous exchange takes place during photosynthesis.	
1.3	The phase of photosynthesis that is independent of light.	
1.4	An energy-rich carbohydrate that is formed during photosynthesis.	
1.5	Photosynthetic tissue in the leaf consisting of elongated cells.	
1.6	Green plants that produce their own food through photosynthesis.	
1.7	Energy from the sun, needed by plants for photosynthesis.	
1.8	Stacks of thylakoids, light dependent phase of photosynthesis takes place here.	
1.9	Green pigment needed for photosynthesis.	
1.10	Part of the chloroplast that contains chlorophyll.	
		(10X1) (10)

Activity 2

Indicate whether each of the statements in COLUMN I apply to A ONLY, B ONLY, BOTH A AND B or NONE of the items in COLUMN II. Write A only, B only, both A and B, or none next to the question number (1.2 to 1.3) in the ANSWER BOOK.

COLUMN I		COLUMN II	
2.1	Molecule that stores energy	A	ATP
		B	ADP
2.2	The organelle in which photosynthesis takes place	A	mitochondria
		B	chloroplast
2.3	Presence of chlorophyll	A	grana
		B	lamella
2.4	Light dependent phase of photosynthesis	A	matrix
		B	stroma
2.5	Gas given off by green plants during photosynthesis	A	O ₂
		B	CO ₂

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno₃ h₂so₄ in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no₂⁺ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0ml of nitric acid and 4.0ml of sulfuric acid and will then react with 4.5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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