

Case Study 35: Emergency Nursing – Chest Trauma

You are the trauma nurse working in the emergency department (ED) of a busy tertiary care facility. You receive a call from the paramedics that they are en route with the victim of gunshot wounds to the chest and abdomen. Perpetrators shot the 32-year-old man during a convenience store robbery. The victim was a customer; the clerk died at the scene. They started 2 16-gauge IVlines with lactated Ringer's solution and oxygen by nonrebreather mask at 15 L/min. The patienthas a sucking chest wound on the left and a wound in the right upper quadrant of the abdomen.

Vital signs in the field are 90/46, 140, and 42. The paramedics state they are having difficulty ventilating the patient, who is diaphoretic, very pale, and lethargic. The estimated time of arrivalis 4 minutes.

1. To help determine the extent of the patient's injuries, the most important question the nurseneeds to ask the paramedics is **a. "How long ago was the patient shot?"** b. "Do you have the weapon that was used?" c. "What was the reason this incident occurred?" d. "Where are the assumed entry and exit wounds?"
2. Describe a sucking chest wound. Air can enter the pleural space through the chest wall duringinspiration due to a penetrating chest wound
3. What are the classic assessment findings associated with a sucking chest wound?
 - Respiratory: Dyspnea, respiratory distress; cough with or without hemoptysis; cyanosisof mouth, face, nail beds, mucous membranes; tracheal deviation, audible air escaping from chest wound; decreased breath sounds; frothy secretions.

Emergency Case Study Scenarios For Nurses

**Megan Kristi Nelson ,Mallory
Philomena Lawrence**



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used

the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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nitration of bromobenzene lab report improved essays 731 words 3 pages open document essay sample check writing quality show more this experiment was performed in order to demonstrate the principles of nitration through electrophilic aromatic substitution it also examined the effects of bromine on an aromatic ring ortho para vs meso directing

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 and H_2SO_4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even closer for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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