

## STUDY GUIDE: Understanding Properties of Matter

1. Use examples to describe the difference between physical and chemical properties of matter. (See handout)
2. List 9 physical properties of matter with an example of each. Density, Ductility, Electrical Conductivity, Magnetism, Malleability, Solubility, Specific Heat, State, Thermal Conductivity (Examples will vary but should be similar to the examples found in the Properties of Matter Booklet.)
3. List 2 chemical properties of matter with an example of each. Flammability, Reactivity (Examples will vary but should be similar to the examples found in the Properties of Matter Booklet.)
4. Which property of matter is observed when water changes from a solid to a liquid? State
5. Adding food coloring to the solution is an example of solubility.
6. If bubbles start to rise from a solution in the scientist's test tube shown, it would be an indication of reactivity.
7. What "point" is observed when water changes from a solid to a liquid? Melting point
8. Which property(ies) would you use to describe most metals? Ductility or Malleability or Electrical Conductivity
9. Which of the properties of metals means that it can be drawn into wires? Ductility
10. Which property allows metals to be bent rather than broken? Malleability
11. Besides being valuable, gold is special because it can be hammered into a very thin foil. This property is called malleability.

Use the table below to interpret questions 12 – 15.

| SUBSTANCE        | DENSITY (g/cm <sup>3</sup> ) | SUBSTANCE      | DENSITY (g/cm <sup>3</sup> ) |
|------------------|------------------------------|----------------|------------------------------|
| Aluminum (solid) | 2.7                          | Ice (solid)    | 0.93                         |
| Pyrite (solid)   | 5.02                         | Water (liquid) | 1.00                         |
| Mercury (liquid) | 13.55                        | Zinc (solid)   | 7.13                         |
| Silver (solid)   | 10.50                        | Wood (oak)     | 0.85                         |

12. What substance has a density more than 13 times greater than water? Mercury
13. What is the density of ice? .93g/cm<sup>3</sup>
14. Which metal is a liquid at room temperature? Mercury
15. What two variables determine the density of a substance? Mass/volume
16. The electromagnets found inside of doorbells are an example of which properties of matter? Magnetism; Electrical conductivity; Ductility (coiled wire around metal)
17. Chemical weathering, such as acid rain causing a statue to fizz and deteriorate is an example of which property of matter? Reactivity
18. Which property of matter is exemplified by pressing metals into coins? Malleability
19. Which property of matter is observed when water changes from a liquid to a solid? State
20. What "point" is observed when water changes from a liquid to a solid? Freezing point
21. Which property of matter is observed when water changes from a liquid to a gas? State
22. What "point" is observed when water changes from a liquid to a gas? Boiling point

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**Lingjun Ying**



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