

Experiment No. 1

Date: / /

Aim: To determine the concentration in terms of molarity of KMnO_4 by titrating it against standard solution of oxalic acid.

Theory: The substances available in the state of high purity are used to prepare standard solutions by dissolving a fixed/definite mass in definite volume of water. Oxalic acid is a primary standard substance. Its molar mass is 126 g/mol

A. Preparation of standard solution of oxalic acid

Molar mass of oxalic acid ($\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$) is 126 g/mol

For 1000 mL 1 M oxalic acid solution required mass is 126 g of oxalic acid

Hence for 100 mL of 0.1 M oxalic acid the required mass is

$$= \frac{126 \times 100 \times 0.1}{1000 \times 1}$$

$$= 1.26 \text{ g of oxalic acid}$$

Apparatus: 100 mL standard flask, balance, watch glass, beaker, glass rod, etc.

Chemicals: Oxalic acid, distilled water.

Procedure:

1. Weigh accurately 1.26 g oxalic acid on watch glass.
2. Transfer the weighed oxalic acid to a beaker and wash the watch glass with distilled water and transfer washings to the beaker. Add little distilled water to dissolve it by stirring.
3. Transfer the solution of oxalic acid from beaker to 100 mL standard flask. Wash beaker twice with water and transfer washings to the 100 mL standard flask. Dilute the solution up to the mark on standard flask to make volume 100 mL.

B. Determination of molarity of KMnO_4 solution using standard solution of oxalic acid.

Apparatus: Burette, pipette, conical flask, burner, water bath, etc.

Chemicals: KMnO_4 solution, 0.1 M oxalic acid, dilute H_2SO_4 .

Procedure:

1. Wash the burette, pipette and conical flask with water.
2. Rinse the burette with given KMnO_4 solution and then fill it. Remove air bubble from the nozzle and avoid leakage if any.
3. Adjust the level of KMnO_4 solution in burette up to zero mark with upper meniscus and fix it on a burette stand.
4. Rinse the pipette with standard 0.1 M oxalic acid solution.
5. Pipette out 10 mL of 0.1 M oxalic acid solution and transfer into clean conical flask. Then add one test tube dilute sulphuric acid. (The solution remains colourless)
6. Heat the conical flask up to 60 to 70°C on wire gauze or water bath (as reaction at room temperature is very slow)
7. Place the conical flask with hot solution on a white porcelain tile below the burette.
8. Start adding KMnO_4 solution drop wise from burette in a conical flask with continuous stirring till light/faint pink colour is obtained in conical flask. Keep the flask constant shaking in a circular manner. This is called swirling of solution. See that the colour does not disappear even on vigorous shaking. This is end point of titration.
9. Note this reading as 'Pilot reading' (pilot reading is always in whole number.)
10. Now fill the burette again with KMnO_4 solution up to the zero mark with upper meniscus.
11. Repeat the above procedure and take minimum three more correct burette readings.
12. Note down the constant burette reading(x) mL (C.B.R.)
13. From C.B.R. calculate the molarity of given KMnO_4 solution.

5

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Observations:

1. Solution in a burette
2. Solution by a pipette
3. Solution in conical flask
4. Indicator
5. End point
6. Chemical Equation

KMnO_4
 oxalic acid
 oxalic acid + 1 tt dil. H_2SO_4
 self
 colourless to faint pink
 $2\text{KMnO}_4 + 5\text{H}_2\text{C}_2\text{O}_4 + 3\text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + 2\text{MnSO}_4 + 8\text{H}_2\text{O} + 10\text{CO}_2\uparrow$

Observation Table:

Burette level	Pilot reading	Burette reading in mL			C.B.R.
		I	II	III	
Final	—	3.5	3.5	3.5	3.5 (x) mL
			0.0	0.0	

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