

CHEMISTRY



Periodic Table of Elements

1 H 1.008																	18 Ar 39.948
2 He 4.0026																	36 Kr 83.80
3 Li 6.941	4 Be 9.0122															54 Xe 131.29	86 Rn 222
5 Na 22.990	6 Mg 24.305															84 Po 209	118 Og 294
9 F 18.998	10 Ne 20.180															85 At 210	117 Ts 294
11 K 39.098	12 Ca 40.078	13 Sc 44.956	14 Ti 47.88	15 V 50.942	16 Cr 52.00	17 Mn 54.938	18 Fe 55.845	19 Co 58.933	20 Ni 58.69	21 Cu 63.546	22 Zn 65.38	23 Ga 69.723	24 Ge 72.64	25 As 74.922	26 Se 78.96	27 Br 79.904	28 Kr 83.80
37 Rb 85.468	38 Sr 87.62	39 Y 88.906	40 Zr 91.224	41 Nb 92.906	42 Mo 95.94	43 Tc 98	44 Ru 101.07	45 Rh 101.07	46 Pd 106.36	47 Ag 107.868	48 Cd 112.411	49 In 114.818	50 Sn 118.710	51 Sb 121.757	52 Te 127.6	53 I 126.905	54 Xe 131.29
55 Cs 132.905	56 Ba 137.327	57 La 138.905	58 Ce 140.12	59 Pr 140.908	60 Nd 144.24	61 Pm 145	62 Sm 150.36	63 Eu 151.964	64 Gd 157.25	65 Tb 158.925	66 Dy 162.50	67 Ho 164.930	68 Er 167.259	69 Tm 168.930	70 Yb 173.054	71 Lu 174.967	86 Rn 222
87 Fr 223	88 Ra 226	89 Ac 227	90 Th 232.038	91 Pa 231.036	92 U 238.029	93 Np 237	94 Pu 244	95 Am 243	96 Cm 247	97 Bk 247	98 Cf 251	99 Es 252	100 Fm 257	101 Md 258	102 No 259	103 Lr 262	118 Og 294

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
 - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatomic:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex:** Central ion + Solvent
 - Transition metal compounds:**
 - Coordination compounds: Non-bonding ligands
 - Organometallics: Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number (Z) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radius:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

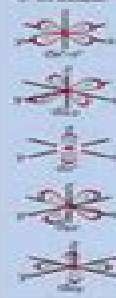
- Missing elements were predicted using periodicity
- Columns 1: Alkali metals
- Columns 2: Alkaline earth metals
- First transition series: Titanium-Copper
- Column 17: Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18: Rare gases: helium, neon, argon, krypton, xenon, and radon
- Actinides: $Z = 90$ or above; radioactive
- Lanthanides: f-subshell elements; starts with $Z = 58$



Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell 5 orbitals



p-subshell 3 orbitals



s-subshell 1 orbital



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge $-q$, $Z = +q$ electrons; for a cation of charge $+q$, $Z = -q$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)



Atomic Quantum Numbers & Orbitals

- n : principle**
 - angular momentum (orbital shape)
 - magnetic (orbital direction)
 - electron spin
- For each n , the possible l values are $0, 1, \dots, n-1$
- For each l , the possible m_l values are $-l, \dots, 0, \dots, +l$
- m_s has two possible values: $+\frac{1}{2}$ and $-\frac{1}{2}$ (spin up, spin down)
- Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell
 - s-orbital: $l = 0$, 1 type
 - d-orbital: $l = 2$, 5 types
 - p-orbital: $l = 1$, 3 types
 - f-orbital: $l = 3$, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



n	l	m_l	m_s	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	0	$-\frac{1}{2}$	N
2	1	+1	$+\frac{1}{2}$	O
2	1	+1	$-\frac{1}{2}$	F
2	1	+1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up, $m_s = +\frac{1}{2}$, and one with spin down, $m_s = -\frac{1}{2}$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
 - First IP: Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

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DP Hallahan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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