

CHAPTER 5

$$3.1 \quad R(t) = \int_0^t C(x) dx = \int_0^t \frac{1}{(x+1)^2} dx = \frac{1}{x+1}$$

$$a) \quad R_0(t) = 1 - \left[1 - \frac{1}{t+1}\right]^2 \quad \rightarrow \quad R_0(5) = 1 - \left[1 - \frac{1}{15}\right]^2 = .96296$$

$$b) \quad 1 - \left[1 - \frac{1}{t+1}\right]^2 = .999$$

$$1 - \frac{1}{t+1} = (1-.999)^{1/2} \quad \rightarrow \quad t = \left[1 - (1-.999)^{1/2}\right]^{-2} - 1 = 1111 \text{ yrs.} = 40.55 \text{ days}$$

$$c) \quad R_0(5) = \frac{1}{15} = .667$$

$$R_0(5) = \sum_{x=0}^5 \left(\frac{1}{x}\right) (.667)^x (.333)^{5-x} = R_0(.667)^5 (.333) + (.667)^6 = .44444 + .296341 = .741185$$

3.3

$$(a) \quad R_0(100) = 1 - \left[1 - e^{-(.0001)(100)}\right]^2 = 2e^{-(.0001)(100)} - e^{-(.0001)(100)} = .99994$$

$$(b) \quad R_0(100) = e^{-(.0001)(100)} - e^{-(.0001)(100)} = .99994 \quad (\text{always reliability})$$

$$3.6 \quad R_0(t) = e^{-\frac{25}{100}\left(\frac{1}{2}\right)^t} = e^{-\frac{1}{4}\left(\frac{1}{2}\right)^t} \quad \rightarrow \quad R_0(1) = e^{-\frac{1}{4}\left(\frac{1}{2}\right)^1} = .99$$

$$-10\left(\frac{1}{2}\right)^8 = .36(.99) \quad \rightarrow \quad R_0 = \left[-\frac{1}{10}(.36(.99))\right]^{-10} = .5588$$

3.7

$$a) \quad R_0 = \left[1 - (1-.97)\left[1 - (1 - (1)(.2)(.95))\right]\right] \left[1 - (1-.97)^2\right] = .985$$

$$b) \quad R_0 = 1 - \left[1 - \left[1 - \left[1 - (.95)(.97)\right]^2\right]\right] \left[1 - (.95)(.95)\right] = 1 - (.005047)(.0075) = .99366$$

$$c) \quad R_0 = \left[1 - (1 - R - R \cdot R)(1 - R)\right] R = \left[1 - (1 - R - R^2 + R^3)\right] R = R^2(1 + R - R^2) = .99$$

R	R_0
.9	.9829
.95	.9936
.97	.9983
.99	.9999

3.8

Introduction To Reliability Maintainability Engineering Solution

Jin-Ying Zhang



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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this

experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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