



# **MARKSCHEME**

**November 2013**

**CHEMISTRY**

**Standard Level**

**Paper 1**

N13 4 Chemi Sp3 Eng Tz0 Xx

**Annelies Wilder-Smith**



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web between the 18th and 19th century mapuche culture and people spread eastwards into the pampas and the patagonian plains this vast new territory allowed mapuche groups to control a substantial part of the salt and cattle trade in the southern cone



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web apr 27 2017 in the native language mapudungun mapuche translates to people of the earth as their name would suggest many mapuche practices and beliefs are deeply connected to the natural world and the mapuche s ancestral land healing ceremonies and the use of mapuche herbal medicine for example are two common cultural practices

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation

$$8 3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8 3 \text{ mole H}_2\text{O}$$

8 3 mole H<sub>2</sub>O  $\times$  18 g H<sub>2</sub>O / 1 mole H<sub>2</sub>O = 149 4 g H<sub>2</sub>O is your theoretical yield so percent yield = 137 1 g H<sub>2</sub>O actual yield / 149 4 g H<sub>2</sub>O  $\times$  100 = 91 77 %

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21 3 grams of O<sub>2</sub> produces a 47 3 % yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$21 3 \text{ grams O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ grams O}_2} = 0 666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = mass of product obtained / calculated mass of the product expected  $\times$  100 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160 0

g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield  
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be  
 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 remains when the reaction is complete

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 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is  
 complete

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 ch<sub>3</sub>oh can be produced by the following reaction 2h<sub>2</sub> co ch<sub>3</sub>oh calculate the theoretical yield of ch<sub>3</sub>oh if 68 5 g of co is  
 reacted with 8 6 g of h<sub>2</sub>

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 which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate  
 what is the

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 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  xs 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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