Engineering Physics Notes: Author: Prayers N Valdys, SDMCET Dharwad...

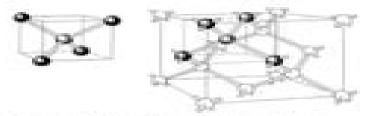


Fig. Tetrahedral up' bonding (very hard!)

The diamond structure has an FCC space lattice and a basis comprising two identical atoms. Alternatively, the structure may be viewed as two interponentating FCC lattices, one displaced relative to the other along a body diagonal by a quarter of its length.

The atoms on four atoms shown imide the main FCC sub-lattice are from other interpenetrated. FCC sub-lattices.

Helatively few elements exist in the diamond structure. Carbon is the most obvious example, but the importance of this structure in solid state physics is mainly due to the fact that this is the natural state of the semiconductors silicon and permanium.

Many two-element compounds form in a diamond-like sinucture in which the first element occupies one FCC sublattice and the second element the other. This structure is named after its prime example, ZuS.

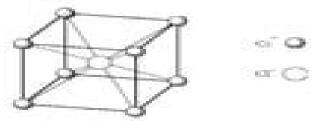
Position of Curbon atoms are at

First Mind: (1900), (1905), (1915), (1915), (1975), (1

Second Rinds (N. N. 10. (N. 16.10) (N. 16.10)

CHOIL

It contains atoms at the corners of a cobe and at the body centre of the cobe. Note that it is NOT a body centered cubic structure (as the atom at the centre is different to the atom at the corners of the cubs).



In the cesium chloride (CsCI) structure, the space fattice is simple cubic. There are two such simple cubic lattices which are separated by one half of the body diagonal of the unit cube. One lattice is occupied by Cs atoms, the other by Cl atoms.

Each atom has 8 nearest neighbors of the opposite kind and is bound to these 8 atoms which are arranged at the comers of the surrounding hexahedron (cube). The bonding is typically issue. The dimensions of the CsCl unit cell are u = b = c = 4.123 Å

Problems:

Obtain the miller indices of a plane which intercepts at a, b/2, 3c in a simple cubic unit cell.

Solution
$$x=a, y=b/2$$
 and $x=3c$

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