



QUESTION 1

1.1.1. Explain the difference between a primary and a secondary cell.

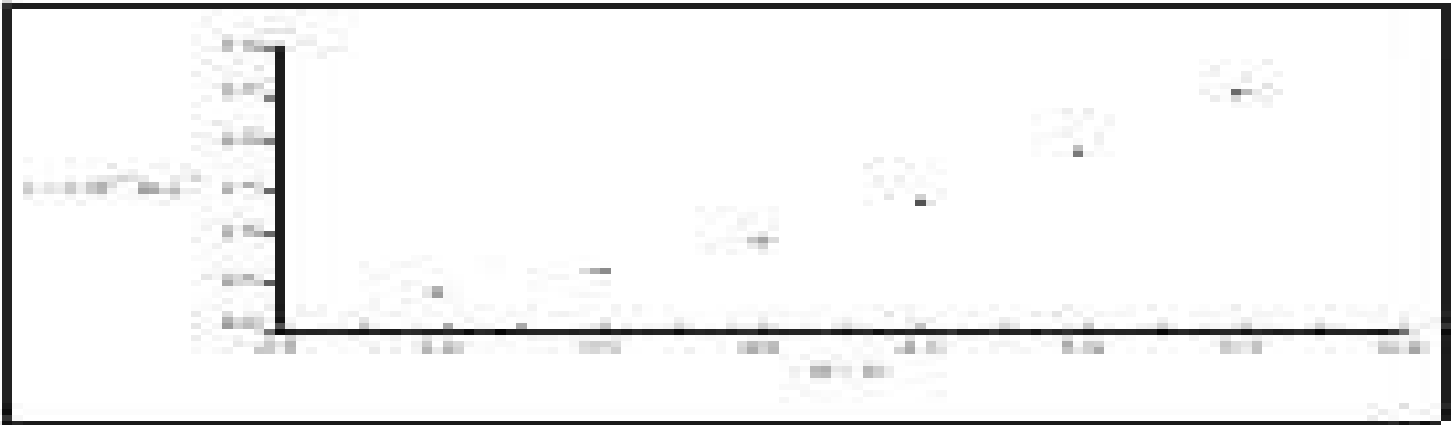
1.1.2. State the half-cell reactions.

1.1.3. The concentration of Zn^{2+} ions in the anode compartment is 0.1 mol dm^{-3} .



1.1.4. When the cell has been connected, it is observed that the concentration of Zn^{2+} ions in the cell decreases. The standard electrode potentials are $E^{\ominus}(Zn^{2+}/Zn) = -0.76\text{ V}$ and $E^{\ominus}(Cu^{2+}/Cu) = +0.34\text{ V}$. Write down the overall cell reaction.

1.1.5. Calculate the cell potential.



1.1.6. Explain the reason for the decrease in the concentration of Zn^{2+} ions.



N12 4 Physi Hp2 Eng Tz0 Xx M

Michael Hoelscher



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