

Questions:

- (1) To identify the correct hydrocarbon structure corresponding to the following formula: $C_{10}H_{18}$
- (2) To calculate the number of moles of the starting material present in the solution.
- (3) To determine the limiting reagent.
- (4) To determine the theoretical amount of product that could be produced.
- (5) To compare the theoretical amount of product with the actual amount of product obtained.

Procedure:

- (1) Put on safety goggles.
- (2) Prepare a solution of the starting material in a beaker.
- (3) Put approximately 10 mL of sodium hydroxide solution in a graduated cylinder.
- (4) Measure the volume of the starting material in the beaker. Pour the sodium hydroxide solution into the beaker.
- (5) Carefully add the graduated cylinder into the beaker with sufficient speed. Stop and approximately 10 mL of sodium hydroxide solution into the graduated cylinder. Repeat several times (indicate your readings). Pour the sodium hydroxide solution back (indicate the resulting measurement your readings) into the graduated cylinder for about one minute. Measure the volume of the solution.
- (6) After adding to the beaker the sodium hydroxide solution, observe the reaction. The mass of the mixture is 10 g and the mass of the solution is your readings.
- (7) Remove a graduated cylinder from the beaker and observe the reaction.
- (8) After the first paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (9) After the second paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (10) After the third paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (11) After the fourth paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (12) After the fifth paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (13) After the sixth paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (14) After the seventh paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (15) After the eighth paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (16) After the ninth paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.
- (17) After the tenth paper with a white solution of sodium hydroxide solution, the contents of the graduated cylinder are the same. Observe the reaction and the mass of the solution. Remove the graduated cylinder from the beaker and observe the reaction. The mass of the solution is your readings.

State your observations:

Observations:	I am writing the answer after completing the experiment. I am writing the answer after completing the experiment.
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Essential Experiments For Chemistry 14d Questions

L Cohen



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