

5 (d) Explain why the current in the external circuit of this cell falls to zero after the cell has operated for some time.

.....

.....

(1 mark)

5 (e) The simplified electrode reactions in a rechargeable lithium cell are

Electrode A       $\text{Li}^+ + \text{MnO}_2 + \text{e}^- \longrightarrow \text{LiMnO}_2$        $E = -0.15 \text{ V}$

Electrode B               $\text{Li}^+ + \text{e}^- \longrightarrow \text{Li}$

Electrode B is the negative electrode.

5 (e) (i) The e.m.f. of this cell is 2.90 V.

Use this information to calculate a value for the electrode potential of electrode B.

.....

.....

(1 mark)

5 (e) (ii) Write an equation for the overall reaction that occurs when this lithium cell is being recharged.

.....

.....

.....

(2 marks)

5 (e) (iii) Suggest why the recharging of a lithium cell may lead to release of carbon dioxide into the atmosphere.

.....

.....

(1 mark)

# Gcse Chemistry Exam Questions And Answers

**Daniel W Foster Professor of Medical  
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## **Table of Contents Gcse Chemistry Exam Questions And Answers**

1. Understanding the eBook Gcse Chemistry Exam Questions And Answers
  - The Rise of Digital Reading Gcse Chemistry Exam Questions And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Chemistry Exam Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gcse Chemistry Exam Questions And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Chemistry Exam Questions And Answers
  - Personalized Recommendations
  - Gcse Chemistry Exam Questions And Answers User Reviews and Ratings

- Gcse Chemistry Exam Questions And Answers and Bestseller Lists
- 5. Accessing Gcse Chemistry Exam Questions And Answers Free and Paid eBooks
  - Gcse Chemistry Exam Questions And Answers Public Domain eBooks
  - Gcse Chemistry Exam Questions And Answers eBook Subscription Services
  - Gcse Chemistry Exam Questions And Answers Budget-Friendly Options
- 6. Navigating Gcse Chemistry Exam Questions And Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Gcse Chemistry Exam Questions And Answers Compatibility with Devices
  - Gcse Chemistry Exam Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gcse Chemistry Exam Questions And Answers
  - Highlighting and Note-Taking Gcse Chemistry Exam Questions And Answers
  - Interactive Elements Gcse Chemistry Exam Questions And Answers
- 8. Staying Engaged with Gcse Chemistry Exam Questions And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gcse Chemistry Exam Questions And Answers
- 9. Balancing eBooks and Physical Books Gcse Chemistry Exam Questions And Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gcse Chemistry Exam Questions And Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gcse Chemistry Exam Questions And Answers
  - Setting Reading Goals Gcse Chemistry Exam Questions And Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gcse Chemistry Exam Questions And Answers
  - Fact-Checking eBook Content of Gcse Chemistry Exam Questions And Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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