

Name _____

Section _____

Date _____

20. A gas sample occupies 200 mL at 700 mm Hg. What volume does the gas occupy at 600 mm Hg?

10.00

$$(200 \text{ mm Hg})(200 \text{ mL}) = (600 \text{ mm Hg}) V_2$$

$$V_2 = 300 \text{ mL}$$

21. Nitrogen gas in a sealed container exerts a pressure of 150 atm at 300 K. What will the pressure be if the tank for 100 mL is left in the sun and the temperature rises to 323 K?

164

$$\frac{150 \text{ atm}}{300 \text{ K}} = \frac{P_2}{323 \text{ K}}$$

$$P_2 = 164 \text{ atm}$$

22. A sample of gas occupies a volume of 80 mL at a pressure of 0.50 atm and a temperature of 273 K. What will be its volume at a pressure of 1.50 atm and a temperature of 323 K?

26.7

$$\frac{(0.50 \text{ atm})(80 \text{ mL})}{273 \text{ K}} = \frac{(1.50 \text{ atm}) V_2}{323 \text{ K}}$$

$$V_2 = 31.55 \text{ mL}$$

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The Princeton Review



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Table of Contents Gas Laws Test Review Answers

1. Understanding the eBook Gas Laws Test Review Answers
 - The Rise of Digital Reading Gas Laws Test Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Test Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Test Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Test Review Answers
 - Personalized Recommendations
 - Gas Laws Test Review Answers User Reviews and Ratings
 - Gas Laws Test Review Answers and Bestseller Lists
5. Accessing Gas Laws Test Review Answers Free and Paid eBooks
 - Gas Laws Test Review Answers Public Domain eBooks
 - Gas Laws Test Review Answers eBook Subscription Services
 - Gas Laws Test Review Answers Budget-Friendly Options
6. Navigating Gas Laws Test Review Answers eBook Formats

- ePub, PDF, MOBI, and More
- Gas Laws Test Review Answers Compatibility with Devices
- Gas Laws Test Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Test Review Answers
 - Highlighting and Note-Taking Gas Laws Test Review Answers
 - Interactive Elements Gas Laws Test Review Answers
- 8. Staying Engaged with Gas Laws Test Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Test Review Answers
- 9. Balancing eBooks and Physical Books Gas Laws Test Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Test Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws Test Review Answers
 - Setting Reading Goals Gas Laws Test Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws Test Review Answers
 - Fact-Checking eBook Content of Gas Laws Test Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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