

## WS Gas Laws - KEY

### Boyle's Law

$$P_1 \times V_1 = P_2 \times V_2 \quad \text{At constant temperature}$$

### Charles's Law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \begin{array}{l} \text{At constant pressure} \\ \text{Temperature in Kelvin} \\ \text{ } ^\circ\text{C} + 273 = \text{K} \end{array}$$

### Lussac's Law

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \begin{array}{l} \text{At constant volume} \\ \text{Temperature in Kelvin} \end{array}$$

### Combined Gas Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

STP: 273 K & 1.00 atm

### Concepts: (circle one choice)

- If the volume of a container of gas is reduced, what will happen to the pressure inside the container?  
increase or decrease
- What happens to the volume of a gas when it is cooled to a lower temperature?  
increase or decrease
- What happens to the pressure in a rigid container when it is slowly warmed?  
increase or decrease

### Sample Calculations:

(Remember: ALL temperatures MUST be in K, NOT °C)

- A sample of gas at 240 K and 670 mmHg occupies a 0.128 L volume. What volume will the gas occupy at 198 K if the pressure remains constant?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(0.128 \text{ L})}{(240 \text{ K})} = \frac{V_2}{(198 \text{ K})} \quad (0.128)(198) = V_2 = \boxed{0.106 \text{ L}}$$

- A sample of gas is in a steel container at  $-75.0^\circ\text{C}$  and 1.48 atm. At what temperature will the sample have a pressure of 7.35 atm?

$$\text{K} = -75 + 273 = 198 \text{ K}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(1.48 \text{ atm})}{(198 \text{ K})} = \frac{(7.35 \text{ atm})}{T_2} \quad T_2(1.48) = (7.35)(198) \quad T_2 = \frac{(7.35)(198)}{(1.48)} = \boxed{983 \text{ K}}$$

- In an airplane, a gas sample occurs at a volume of 1.50 L at 760 mmHg. Suppose, while flying, the airplane loses pressure and the volume of the gas increases to 11.40 L. What is the pressure in the airplane if the temperature is constant?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad (760 \text{ mmHg})(1.50 \text{ L}) = P_2(11.40 \text{ L}) \quad (760)(1.50) = P_2 = \boxed{100 \text{ mmHg}}$$

# Gas Law Simulation Lab Answer Key

**SJ Ball**



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