

# Lab Data

Report pressure of water (mmHg) 21.1

Verify your calculation. In your set-up, do your units cancel?

Did you report your data to the correct number of significant figures?

Volume of hydrogen gas collected (mL) 15.00

Volume of hydrogen gas collected (L) .01500

Pressure of hydrogen gas (mmHg) 759.9

Pressure of hydrogen gas (atm)

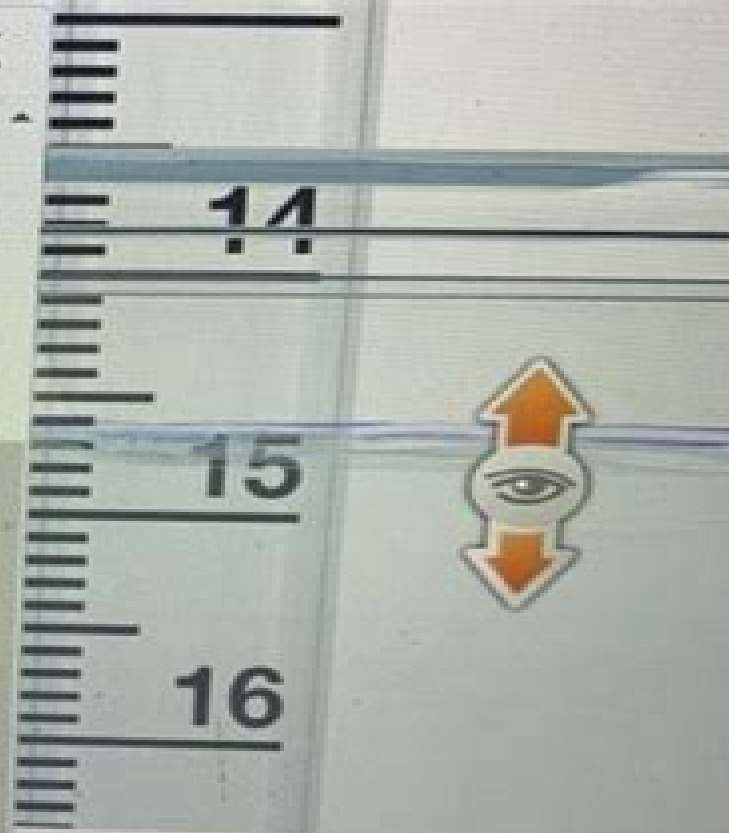
Moles of hydrogen gas produced (mol)

Ideal gas constant ( $\text{L}\cdot\text{atm}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ )

How to calculate ideal gas constant

Water Vapor Pressure Table

X



MY NOTES

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Sin A  $\frac{a}{c} = \sin B \frac{b}{c}$   
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# Gas Law Constant Chemistry Lab Report Answers

**J Dewey**



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