

**Skill - Student will accurately describe the relationship between P, V, T and n of gases.**

1. if the temperature increases, the volume will...

**INCREASE,** an inc in temp causes an inc in volume due to a direct relationship.

2. if the pressure decreases the volume will...

**INCREASE,** a dec in pressure will cause an inc in volume due to an inverse relationship.

3. if the moles of gas decrease the volume will...

**DECREASE,** a dec in moles will cause a dec in volume due to a direct relationship.

**Skill - Student will accurately calculate a 2-variable gas law problem.**

4. A sample of methane gas that has a volume of 10.0 L at 25°C is cooled to 5°C at a constant pressure. Calculate the new volume. What equation must you use? **SHOW YOUR WORK.**

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \rightarrow \frac{10.0\text{L}}{298\text{K}} = \frac{V_2}{278\text{K}}$$

5. A gas occupies 12.3 liters at a pressure of 40.0 mmHg. What is the volume when the pressure is increased to 60.0 mmHg?

$$P_1V_1 = P_2V_2 \rightarrow (40\text{ mmHg})(12.3\text{L}) = (60\text{mmHg})(V_2) \quad \mathbf{V_2 = 8.2\text{L}}$$

**Skill - Student will accurately calculate using the ideal gas law.**

6. Radon, a radioactive gas formed naturally in the soil, can cause lung cancer. It can pose a hazard to humans by seeping in to houses and become trapped in basements. A 1.5 mol sample of radon has a volume of 23.0 L at 33°C. What is the pressure of the gas? What equation must you use? **SHOW YOUR WORK.**

$$PV = nRT \rightarrow P = nRT/V \\ = (1.5\text{mol} \times 0.08206\text{L}\cdot\text{atm}/\text{K}\cdot\text{mol} \times 306\text{K})/23.0\text{L} \rightarrow \mathbf{P = 1.8\text{atm}}$$

7. A 3.25g sample of solid calcium carbide ( $\text{CaC}_2$ ) reacts with water to produce acetylene gas ( $\text{C}_2\text{H}_2$ ) and aqueous calcium hydroxide. If the temperature and pressure within the system are 17°C and 740.0 mmHg, how many milliliters of acetylene were produced? **SHOW YOUR WORK.**



$$3.25\text{g CaC}_2 \times (1\text{mol CaC}_2/64.1\text{g CaC}_2) \times (1\text{mol C}_2\text{H}_2/1\text{mol CaC}_2) = 0.051\text{mol C}_2\text{H}_2$$

$$PV = nRT \rightarrow V = nRT/P = (0.051\text{mol} \times 0.08206\text{L}\cdot\text{atm}/\text{K}\cdot\text{mol} \times 290\text{K})/0.974\text{atm} = 1.24\text{L} = \mathbf{1240\text{mL}}$$

# Gas Laws Review Answer Key

**LL Leslie**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LL Leslie' text.

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