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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	kJ mol ⁻¹	kJ mol ⁻¹
bond energy in kJ/mol	440	435

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of one N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (500) + 3(440) \\
 & 440 + 3(440) \quad \quad \quad 1660 \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -90 \\
 & 1660 - x = -90 \\
 & x = 1660 + 90 \\
 & \quad \quad \quad 1750 \text{ kJ}
 \end{aligned}$$
- Calculate the energy of one N-H bond.

$$\begin{aligned}
 & 6 \times (\text{N-H}) = 1750 \\
 & \text{N-H} = \frac{1750}{6} \quad \quad \quad 292 \text{ kJ/mol}
 \end{aligned}$$

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