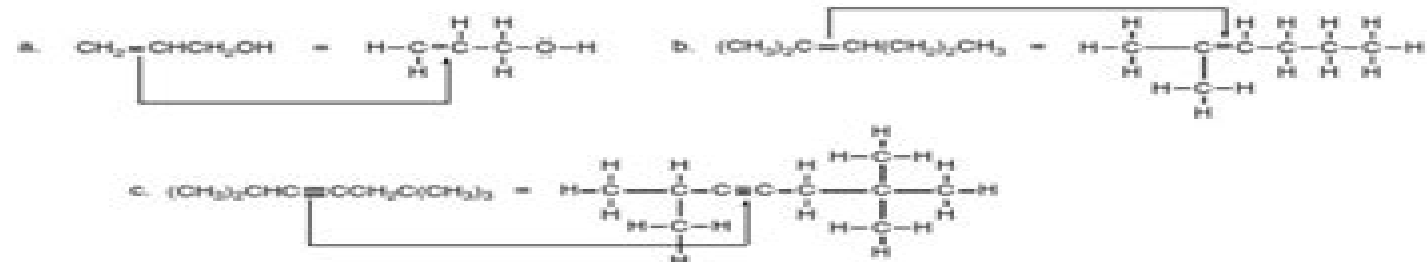# Chapter 13 Unsaturated Hydrocarbons

Solutions to In-Chapter Problems

13.1   To draw a complete structure for each condensed structure, first draw in the multiple bonds. Then draw in all the other C's and H's, as in Example 13.1.

13.2   To determine whether each molecular formula corresponds to a saturated hydrocarbon, alkene, or alkyne, recall that the formula for a saturated hydrocarbon is C_nH_{2n+2}, the formula for an alkene is C_nH_{2n}, and the formula for an alkyne is C_nH_{2n-2}.

a. $C_3H_6 = C_nH_{2n} =$ alkene
b. $C_5H_{12} = C_nH_{2n+2} =$ saturated hydrocarbon
c. $C_6H_{14} = C_nH_{2n-2} =$ alkyne
d. $C_6H_{12} = C_nH_{2n} =$ alkene

13.3   Use the general formulas [saturated hydrocarbon (C_nH_{2n+2}), alkene (C_nH_{2n}), and alkyne (C_nH_{2n-2})] to determine the molecular formula for each compound.

a. alkene = C_nH_{2n}, $4 \times 2 = 8$, C_4H_8
b. saturated hydrocarbon = C_nH_{2n+2}, $(6 \times 2) + 2 = 14$, C_6H_{14}
c. alkyne = C_nH_{2n-2}, $(7 \times 2) - 2 = 12$, C_7H_{12}
d. alkene = C_nH_{2n}, $5 \times 2 = 10$, C_5H_{10}

13.4   Give the IUPAC name for each compound using the following steps, as in Example 13.2:
[1] Find the longest chain containing both carbon atoms of the multiple bond.
[2] Number the chain to give the double bond the lower number.
[3] Name and number the substituents and write the complete name.

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