

1. What might happen if you mixed a strong acid with an equally strong base?

- | | | | |
|-------------------------|----------------------------------|-------------------------|---|
| ☐ A | The acid would destroy the base. | ☐ B | You would see an explosive chemical reaction. |
| ☐ C | The base would destroy the acid. | ☐ D | You'd wind up with a pH-neutral substance. |

2. $\text{PO}_4^{3-} + \text{HNO}_3 \rightarrow \text{NO}_3^- + \text{HPO}_4^{2-}$ What is the conjugate acid in the following equation?

- | | | | |
|-------------------------|---------------------|-------------------------|--------------------|
| ☐ A | HPO_4^{2-} | ☐ B | PO_4^{3-} |
| ☐ C | HNO_3 | ☐ D | NO_3^- |

3. $\text{HCO}_3^- + \text{HCl} \rightarrow \text{H}_2\text{CO}_3 + \text{Cl}^-$ What is the conjugate base in the following reaction?

- | | | | |
|-------------------------|-------------------------|-------------------------|------------------|
| ☐ A | Cl^- | ☐ B | HCl |
| ☐ C | H_2CO_3 | ☐ D | HCO_3^- |

4. $\text{HSO}_4^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$
Identify the Acid in the above reaction.

- | | | | |
|-------------------------|--------------------|-------------------------|------------------------|
| ☐ A | SO_4^{2-} | ☐ B | H_3O^+ |
| ☐ C | HSO_4^- | ☐ D | H_2O |

5. $\text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq})$
Using the above reaction, which compound is the conjugate base?

- | | | | |
|-------------------------|----------------------|-------------------------|--------------------|
| ☐ A | HCO_3^- | ☐ B | OH^- |
| ☐ C | H_2O | ☐ D | CO_3^{2-} |

6. What are the products to a neutralization reaction?

- | | | | |
|-------------------------|----------------------------|-------------------------|-----------------------------------|
| ☐ A | OH^- + Ionic Salt | ☐ B | H_2O + Ionic Salt |
|-------------------------|----------------------------|-------------------------|-----------------------------------|

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1

Didier Musso



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