

Intro to Chemical Reactions:

Test Review Answer Key



1. What do the symbols (s), (l), (g), and (aq) represent? (s) = solid, (l) = liquid, (g) = gas, and (aq) = aqueous
2. Why must chemical equations be balanced? Chemical equations must be balanced to satisfy the Law of Conservation of Mass. The number of atoms of the reactant side must equal the number of atoms on the product side of a chemical equation.
3. How can we be certain that a chemical reaction has taken place? Color change, temperature change, formation of a precipitate, production of gas (bubbles), and light emission.
4. Balance the following chemical equations:
 - a. $1 \text{ AsCl}_3 + 3 \text{ H}_2\text{O} \rightarrow 3 \text{ HCl} + 1 \text{ As(OH)}_3$
 - b. $4 \text{ Na} + 1 \text{ O}_2 \rightarrow 2 \text{ Na}_2\text{O}$
 - c. $1 \text{ Cu} + 2 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ CuSO}_4 + 1 \text{ SO}_2 + 2 \text{ H}_2\text{O}$
 - d. $2 \text{ Sb} + 3 \text{ H}_2\text{O} \rightarrow 1 \text{ Sb}_2\text{O}_3 + 3 \text{ H}_2$
 - e. $1 \text{ La(NO}_3)_3 + 3 \text{ NaOH} \rightarrow 1 \text{ La(OH)}_3 + 3 \text{ NaNO}_3$
 - f. $1 \text{ Ca(AlO}_2)_2 + 8 \text{ HCl} \rightarrow 2 \text{ AlCl}_3 + 1 \text{ CaCl}_2 + 4 \text{ H}_2\text{O}$
 - g. $1 \text{ C}_4\text{H}_8 + 6 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
 - h. $2 \text{ C}_3\text{H}_8\text{O}_3 + 7 \text{ O}_2 \rightarrow 6 \text{ CO}_2 + 8 \text{ H}_2\text{O}$
5. Identify the following types of chemical reactions:
 - a. $2 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$ **Synthesis or Combustion**
 - b. $2 \text{ H}_2\text{O} \rightarrow 2 \text{ H}_2 + \text{O}_2$ **Decomposition**
 - c. $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$ **Single Replacement**
 - d. $2 \text{ CO} + \text{O}_2 \rightarrow 2 \text{ CO}_2$ **Synthesis or Combustion**
 - e. $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$ **Decomposition**
 - f. $2 \text{ KBr} + \text{Cl}_2 \rightarrow 2 \text{ KCl} + \text{Br}_2$ **Single Replacement**
 - g. $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$ **Synthesis**
 - h. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$ **Double Replacement**
 - i. $2 \text{ H}_2\text{O}_2 \rightarrow 2 \text{ H}_2\text{O} + \text{O}_2$ **Decomposition**
 - j. $\text{Ca(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + 2 \text{ H}_2\text{O}$ **Double Replacement**

Intro To Chemical Reactions Test

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Intro To Chemical Reactions Test:

Intro To Chemical Reactions Test Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Intro To Chemical Reactions Test**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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