

## Chapter 2

### *The First Law of Thermodynamics*

2.2. Work is defined as *negative*  $p\Delta V$  because if a system does work on the surroundings, the system *loses* energy.

$$2.4. \quad 1780 \text{ torr} \left| \frac{1 \text{ atm}}{(760 \text{ torr})} \right| = 2.34 \text{ atm}$$

$$w = -P\Delta V = -2.34 \text{ atm} (1.00 \text{ L} - 3.55 \text{ L}) \left( \frac{101.32 \text{ J}}{1 \text{ L} \cdot \text{atm}} \right) = +605 \text{ J}$$

605 J of work are done ON the system.

- 2.6. (a) The work would be less because the external pressure is less.  
(b) The work would be greater because the external pressure is greater.  
(c) No work would be performed because the external pressure is (effectively) zero.

2.8. First, we need to find the final volume of the  $\text{CO}_2$ . Using  $P_1V_1 = P_2V_2$ , we find that  $V_2 = 105 \text{ mL}$ .

$$w = -(1.0 \text{ atm})(105 - 25.0 \text{ mL}) \left( \frac{1 \text{ L}}{1000 \text{ mL}} \right) \left( \frac{101.32 \text{ J}}{1 \text{ L} \cdot \text{atm}} \right) = -8.1 \text{ J}$$

2.10. First, determine the  $\Delta T$ :  $330 \text{ K} - 298 \text{ K} = 32 \text{ K}$ .

$$\text{Using } q = m \cdot c \cdot \Delta T, \text{ we rearrange: } c = \frac{q}{m \cdot \Delta T}$$

$$\text{Substituting: } c = \frac{288 \text{ J}}{(50.5 \text{ g})(32 \text{ K})} = 0.178 \frac{\text{J}}{\text{g} \cdot \text{K}}$$

2.12. First, we need to calculate the number of moles of phosphorous:

$$\left( \quad \right)$$

# Gases Solution Manual

**SA Dillow**



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