

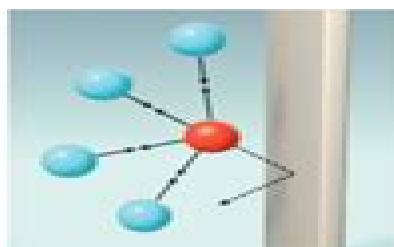
# Chapter 11: Gases

**Focus of this section:**

**Deviation from ideality**

## Deviation from Ideal Behavior:

- recall: real gases deviate from ideal gas behavior at high pressure and low temperature conditions
  1. real molecules have volume
    - cannot compress gas to zero volume
  2. real molecules interact (attracted to each other)
    - causes lower  $P$  (vs. ideal gas), especially at lower temperatures



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$$= P_{\text{obs}} < P_{\text{ideal}}$$

= need to add correction to pressure term

= Since 2 molecules cannot occupy the same space,  $V_{\text{accessible}} < V_{\text{container}}$  and a corrective term should be subtracted from  $V$  in ideal gas equation

$$\left( \underbrace{P + \frac{an^2}{V^2}}_{\text{corrected pressure term}} \right) \left( \underbrace{V - nb}_{\text{corrected volume term}} \right) = nRT$$

$PV = nRT$

$$P_{\text{ideal}} = \left( P_{\text{obs}} + \frac{n^2 a}{V^2} \right)$$

$$\text{correction} \propto \left( \frac{n}{V} \right)^2$$

$$V_{\text{ideal}} = (V_{\text{container}} - nb)$$

$$\text{correction} \propto n$$

= Note, "a" and "b" are different for each gas

# Gases Section Quizzes

**Patrick Vollmar**



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