

# CHAPTER 10: GASES

## Exercises

### PROBLEM SET 10.1

**PROBLEM 10.1.1** Answer the following questions in the space provided.

1.     a     Pressure is force per unit area. What units do we use to measure pressure? (List all units.)  
 (a) Pascals  
 (b) Atmospheres  
 (c) mmHg  
 (d) Torr

2.     a, b, c, d     Show the following pressure in the missing units:

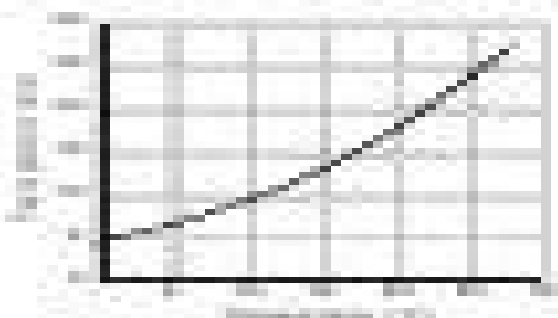
1 atm = 101.3 kPa      1 atm = 101.3 kPa  
 1 atm = 101.3 kPa      1 atm = 101.3 kPa

3. Suppose that molecules of a gas are confined in a cylinder with a piston. The gas is at a constant temperature.

Sketch and label a graph of pressure vs. volume at a constant temperature. From this graph, explain the relationship between pressure and volume of a gas.

**PROBLEM 10.1.2** Answer the question in the space provided. Show all work in the space provided.

4. a. Suppose we have a gas at a constant temperature. If the volume of the gas is doubled, the pressure will be halved. Explain why this is true using the kinetic theory of gases.



- b. Sketch a graph of pressure vs. volume at a constant temperature.
- c. Sketch a graph of pressure vs. temperature at a constant volume.

# Gases Section 1 Review Answers

**M Walker**



## **Gases Section 1 Review Answers:**

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