

CHEM 101 Worksheet, Chapter 5: Gases

Multiple choice questions 1-10: circle the correct answer(A-E)

Show your work questions 11-14: fill in the answer

1 atm = 760 mm Hg = 101,325 Pa 0 °C = 273 K R = 0.08206 L.atm/(mol.K) = 8.314 J/(mol.K)

1. 2.67×10^5 Pa = _____ mm Hg?

- A. 3.56×10^6 mm Hg B. 3.47×10^{-2} mm Hg C. 3.51 mm Hg D. 20.0 mm Hg E. 200 mm Hg

2. How many pascal (Pa) is 7.5×10^{-2} mm Hg?

- A. 100 Pa B. 5.78×10^4 Pa C. 10.0 Pa D. 9.97×10^{-7} Pa E. 1000 Pa

3. What is the volume occupied by 3.36 g $\text{NH}_3(\text{g})$ when the pressure is 740 mm Hg and its temperature 27 °C?

- A. 5.00 L B. 4.74 L C. 1.44×10^3 L D. 0.117 L E. 500 L

4. What is the mass of CO_2 in a 340. mL container when its pressure is 2.80 atm and the temperature is 57 °C?

- A. 26.7 g B. 0.125 g C. 125 g D. 3.52 g E. 3.47×10^{-4} g

5. A certain amount of gas at 300K is held in a volume of 1.0L. When the gas is heated to 600K (the volume is still 1.0 L) the pressure will:

- A. stay the same B. be four times as high C. be double the initial pressure
D. be one half the initial pressure E. be one fourth of the initial pressure

6. Which of the following gases will have the highest density, when all are at the same temperature and pressure?

- A. $\text{CO}_2(\text{g})$ B. $\text{NO}_2(\text{g})$ C. $\text{C}_3\text{H}_8(\text{g})$ D. $\text{SO}_2(\text{g})$ E. $\text{N}_2\text{O}_4(\text{g})$

7. Calculate the density of butane, $\text{C}_4\text{H}_{10}(\text{g})$ at p = 800 mm Hg and T = 37 °C.

- A. 2.17 g/L B. 0.272 g/L C. 2.17 g/m³ D. E. 0.0236 g/L E. 2.40 g/L

8. An unknown gas has a density of 2.67 g/L at p = 1.0 atm and T = 19 °C. What is the molar mass of the gas?

- A. 64.0 g/mol B. 5.76 g/L C. 640 g/mol D. 64. 0 kg/mol E. 32.0 g.mol

9. What volume of $\text{CO}_2(\text{g})$ will be formed when 2.0 g $\text{CaCO}_3(\text{s})$ is heated and the gas is collected at 700 mm pressure and 20 °C? The reaction is: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$

- A. 0.44 L B. 0.52 L C. 52 L D. 3.61 L E. 22.4 L

10. A mixture of 3.0 mol $\text{H}_2(\text{g})$ and 2.0 mol $\text{N}_2(\text{g})$ has a total pressure of 4.5 atm. What are the partial pressures of H_2 and N_2 in the mixture?

- A. 2.5 atm H_2 , 2.0 atm N_2 B. 0.5 atm H_2 , 4.0 atm N_2 C. 1.8 atm H_2 , 2.7 atm N_2
D. 4.0 atm H_2 , 0.5 atm N_2 E. 2.7 atm H_2 , 1.8 atm N_2

Continued next page

Gases Multiple Choice Questions

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Table of Contents Gases Multiple Choice Questions

1. Understanding the eBook Gases Multiple Choice Questions
 - The Rise of Digital Reading Gases Multiple Choice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Multiple Choice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Multiple Choice Questions
 - Personalized Recommendations

- Gases Multiple Choice Questions User Reviews and Ratings
- Gases Multiple Choice Questions and Bestseller Lists
- 5. Accessing Gases Multiple Choice Questions Free and Paid eBooks
 - Gases Multiple Choice Questions Public Domain eBooks
 - Gases Multiple Choice Questions eBook Subscription Services
 - Gases Multiple Choice Questions Budget-Friendly Options
- 6. Navigating Gases Multiple Choice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases Multiple Choice Questions Compatibility with Devices
 - Gases Multiple Choice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases Multiple Choice Questions
 - Highlighting and Note-Taking Gases Multiple Choice Questions
 - Interactive Elements Gases Multiple Choice Questions
- 8. Staying Engaged with Gases Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Gases Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Multiple Choice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gases Multiple Choice Questions
 - Setting Reading Goals Gases Multiple Choice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gases Multiple Choice Questions
 - Fact-Checking eBook Content of Gases Multiple Choice Questions

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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