

7. If you have 17 liters of gas at a temperature of 67 °C and a pressure of 88.89 atm, what will be the pressure of the gas if you raise the temperature to 94 °C and decrease the volume to 12 liters?

$$P_2 = \frac{P_1 V_1 T_2}{V_2 T_1} = \frac{88.89 \text{ atm} (17 \text{ L}) 367 \text{ K}}{(12 \text{ L}) 340 \text{ K}}$$

$$P_2 = \quad \text{atm}$$

8. I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{1.2 \text{ atm} (48 \text{ L}) 325 \text{ K}}{320 \text{ K} (0.5 \text{ atm})}$$

$$V_1 = \quad \text{L}$$

9. If you have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of gas if you decrease the pressure to 45 atm and decrease the temperature to 750 K?

$$V_2 = \frac{78 \text{ atm} (21 \text{ L}) 750 \text{ K}}{900 \text{ K} (45 \text{ atm})}$$

$$V_2 = \quad \text{L}$$

10. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50 °C, what will the temperature of the gas be if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

$$1/T_2 = \frac{P_1 V_1}{T_1 P_2 V_2} = \frac{(2.9 \text{ L}) 5 \text{ atm}}{323 \text{ K} (3 \text{ atm}) 2.4 \text{ L}}$$

$$T_2 = 160 \text{ K}$$

11. You have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 L, how many liters of gas did you start with?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{115 \text{ K} (29 \text{ L}) 30 \text{ atm}}{(225 \text{ K}) 60 \text{ atm}}$$

$$V_1 = \quad \text{L}$$

Gas Laws Practice Calculations Answers

TD Snyder



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Table of Contents Gas Laws Practice Calculations Answers

1. Understanding the eBook Gas Laws Practice Calculations Answers
 - The Rise of Digital Reading Gas Laws Practice Calculations Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Calculations Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Practice Calculations Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Calculations Answers
 - Personalized Recommendations
 - Gas Laws Practice Calculations Answers User Reviews and Ratings
 - Gas Laws Practice Calculations Answers and Bestseller Lists
5. Accessing Gas Laws Practice Calculations Answers Free and Paid eBooks
 - Gas Laws Practice Calculations Answers Public Domain eBooks
 - Gas Laws Practice Calculations Answers eBook Subscription Services
 - Gas Laws Practice Calculations Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Gas Laws Practice Calculations Answers Compatibility with Devices
 - Gas Laws Practice Calculations Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Practice Calculations Answers
 - Highlighting and Note-Taking Gas Laws Practice Calculations Answers
 - Interactive Elements Gas Laws Practice Calculations Answers
8. Staying Engaged with Gas Laws Practice Calculations Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Practice Calculations Answers
9. Balancing eBooks and Physical Books Gas Laws Practice Calculations Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Practice Calculations Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Practice Calculations Answers
 - Setting Reading Goals Gas Laws Practice Calculations Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws Practice Calculations Answers
 - Fact-Checking eBook Content of Gas Laws Practice Calculations Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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